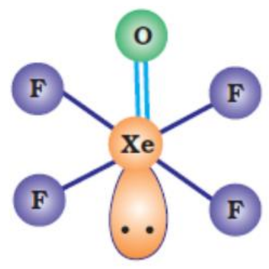
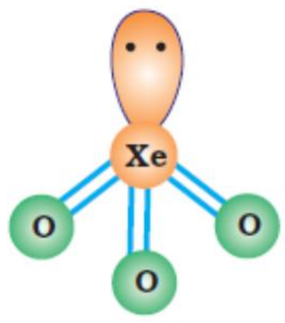
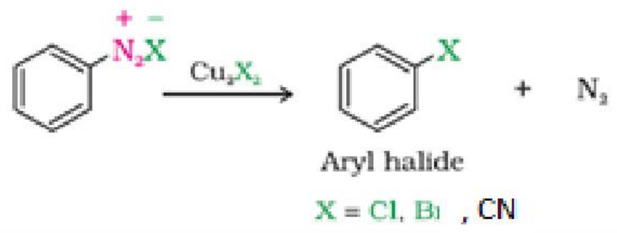
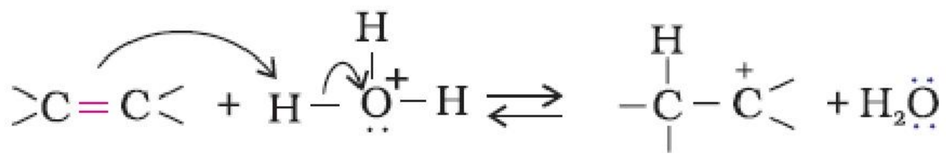
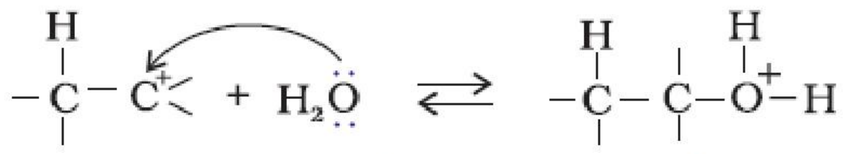
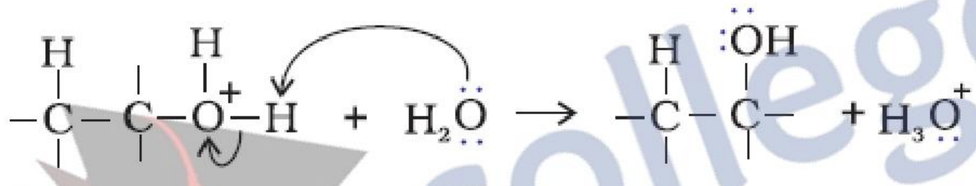
Marking scheme – 2017 (Compartment)

CHEMISTRY (043)/ CLASS XII

Set 56/1/3

Q.No	Value Points	Marks
1	Having no α- hydrogen	1
2	Frenkel Defect	1
3	$K_4[Fe(CN)_6]$ / $[Fe(CN)_6]^{4-}$	1
4	Orbital splitting energies are not sufficiently large for forcing pairing	1
5	2,3-dinitro phenol	1
6	(i) First order (ii) s^{-1} / time $^{-1}$	1 1
7	i) In NH_4^+, all are bond pairs whereas in ammonia the lone pair of electron on nitrogen repels the bond pairs and reduces the bond angle. ii) I-Cl bond is weaker than I-I bond / low bond dissociation enthalpy in I-Cl	1 1
8	Vapour pressure of the solvent decreases in the presence of non – voilatile solute (glucose) hence boiling point increases	2
9	i) (b) is chiral ii) (a) will undergo S_N2 reaction faster	1 1
10.	Hypophosphorous acid is a good reducing agent as it contains two P-H bonds. There is no P-H bond in orthophosphoric acid , so it is not a reducing agent Example : It reduces $AgNO_3$ to metallic silver/ chemical equation	1 1
	OR	
10	a) 4 b) Due to lower bond dissociation enthalpy of BiH_3 as compared to SbH_3	1 1
11	i) The process of removing an adsorbed substance from a surface on which it is adsorbed. ii) The formation of micelles takes place only above a particular concentration called CMC. iii) The catalytic reaction that depends upon the pore structure of the catalyst and size of the reactant and product molecules.	1 1 1
12	a) $H_2O < H_2S < H_2Se < H_2Te$, because of decrease in bond dissociation enthalpy. b) 	1,1 1
	OR	
12	a) i)Due to higher oxidation state of P in PCl_5 ii) Liberation of hydrogen prevents the formation of $FeCl_3$ b) 	1 1 1

*These answers are meant to be used by evaluators

13	a) A: CH_3- $CH{=}CH_2$ B: CH_3- CH_2-CH_2Br C: CH_3- CH_2-CH_2I D: CH_3- CH_2-CH_2MgI  b)	½ × 4 1
14	a) CH_3-O-CH_3 + HI ⟶ CH_3-OH + CH_3-I b) . Protonation of alkene to form carbocation by electrophilic attack of H_3O^+. $H_2O + H^+ \rightarrow H_3O^+$  Nucleophilic attack of water on carbocation.  Deprotonation to form an alcohol.  	1 ½ ½ 1
15	In bcc, z=2 ; d = (zxM) /a^3x N_A (i) No. of atoms = $\frac{w}{M}$ × N_A $2.5 \times 10^{24} = \frac{250\,g}{M} \times N_A$ M= $250 \times NA / 2.5 \times 10^{24}$ (ii) Putting values of M in equation (i) d= 2 x 250 g × N_A / [2.5×10^{24} atoms × $(400 \times 10^{-10}$ cm$)^3$ x N_A] d= 3.125 g/ cm^3 (or any other correct method)	½ 1 ½ 1
16	$P_A = 2Po - Pt$ $= (2 \times 35) - 63 = 7$ k= $\frac{2.303}{t}$ log Po/P_A k= $\frac{2.303}{100}$ log 35/7 k= $\frac{2.303}{100} \times 0.6990$ $= 2.236 \times 10^{-3}\ s^{-1}$ (or any other correct method)	½ ½ 1 1
17	i) A: CH_3- CH_2CN ; B: CH_3- CH_2- CH_2NH_2 ; C: CH_3- CH_2- CH_2-NH-$COCH_3$ ii) A: $Ar-\overset{+}{N_2}B\overset{-}{F_4}$; B:  C :	½ ×3 ½ ×3
18.	a) Glycosidic linkage b) Source : Meat, Fish, egg, curd (any one) ; Pernicious anaemia	1

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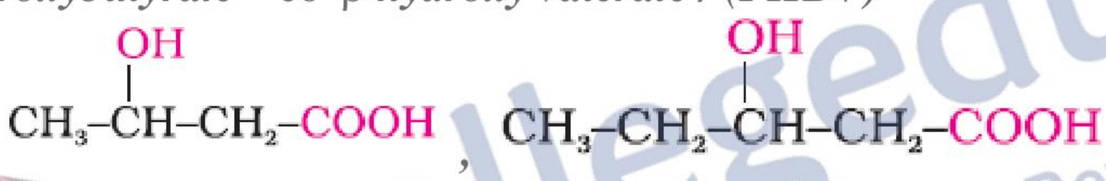
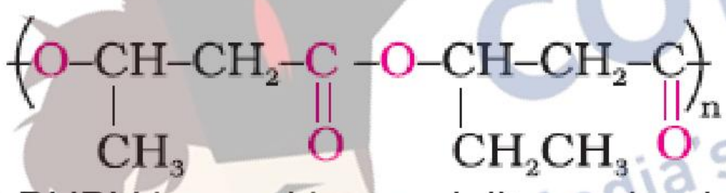
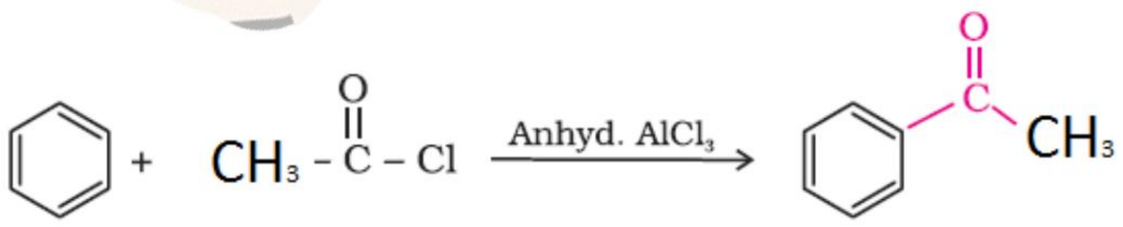
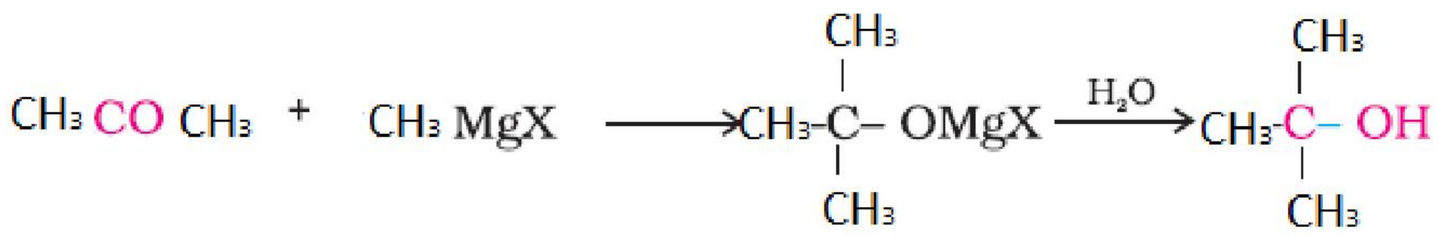
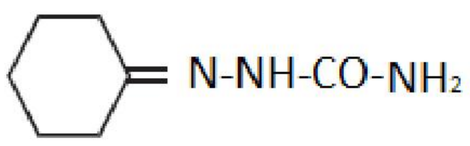
	c) DNA is double strand while RNA is single strand molecule (or any other correct difference)	½ , ½ 1
19.	Hybridisation : dsp^2 Magnetic character : Diamagnetic Spin nature: Low spin	1 1 1
20.	i) Controlling depression and hypertension Class : Tranquilizers ii) Relieve pain and reduce fever Class: Non- Narcotic analgesics / Analgesics iii) Kills or inhibits the growth of micro organisms Class: Antibiotics	½ ½ ½ ½ ½ ½
21	$p_{total} = p_1^\circ + (p_2^\circ - p_1^\circ)\ x_2$ $600 = 450 + (700-450)\ x_2$ $x_2 = 0.6$ $x_2 = 1 - 0.6 = 0.4$	1 1 ½ ½
22	a) Impurities are more soluble in the melt than in the solid state of the metal. Example : Ge/ Si/ B (any other) b) i)Zn/ Hg ii) Sn	½ ½ 1 1
23	a) ***Poly β-hydroxybutyrate – co-β-hydroxy valerate / (PHBV)*** ***Monomers :***  ***Repeating unit :***  *b)* PHBV is used in speciality packaging, orthopaedic devices and in controlled release of drugs.(any two) c) Concern for environment , caring (or any other)	½ ½ , ½ ½ ½ , ½ ½ , ½
24	a) i).  ii) .  b) i) Because it is a deactivating group / Due to electron withdrawing carboxylic group resulting in decreased electron density at o- and p- position. ii) Due to extensive association of carboxylic acid molecules through intermolecular hydrogen bonding. iii) Due to steric and + I effect of two methyl groups in propanone	1 1 1 1 1
	OR	
24	a) i) . 	1

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	ii) CH_3COOH iii) CH_3 -CH(Br)-COOH b) i) Add ammonical solution of silver nitrate / Tollen's reagent to both the compounds, propanal will give silver mirror while propanone does not. ii) Add $NaHCO_3$ solution to both the compounds, Benzoic acid will give effervescence and liberate CO_2 while benzaldehyde will not. (Or any other suitable test)	1 1 1 1
25	a) E^0 value of silver is lower than that of gold, hence silver displaces gold which gets deposited on the silver object. E^0 value of copper is lower than that of silver, hence silver cannot displace copper from its solution. b) i) Electrons flow from Zn to Ag plate. ii) Zn as anode and Ag acts as cathode iii) Cell will stop functioning iv) Concentration of Zn^{2+} ions will increase and that of Ag^+ ions will decrease. v) No change	1 1 ½ ½ ½ ½, ½ ½
	OR	
25	a) When concentration approaches zero, the molar conductivity is known as **limiting molar conductivity** The change in Λm with dilution is due to the increase in the degree of dissociation and consequently the number of ions in the total volume of the solution that contains 1 mol of electrolyte , hence Λm increases steeply. b) $E_{cell} = E^o_{cell} - \frac{0.059}{n} \log \frac{[Mg2+]}{[Cu2+]}$ $=2.71\ V - \frac{0.059}{2} \log \frac{0.1}{0.001}$ $=2.71\ V - \frac{0.059}{2} \log 10^2$ $= 2.651\ V$	1 1 1 1
26	a) A: Na_2CrO_4 ; B: $Na_2Cr_2O_7$; C : $K_2Cr_2O_7$ $4\ FeCr_2O_4 + 8\ Na_2CO_3 + 7\ O_2 \rightarrow 8\ Na_2CrO_4 + 2\ Fe_2O_3 + 8\ CO_2$ $2Na_2CrO_4 + 2\ H^+ \rightarrow Na_2Cr_2O_7 + 2\ Na^+ + H_2O$ $Na_2Cr_2O_7 + 2\ KCl \rightarrow K_2Cr_2O_7 + 2\ NaCl$	½ , ½ , 1 1 1 1
	OR	
26	a) i)Copper; Due to high $\Delta_a H^\ominus$ and low $\Delta_{hyd} H^\ominus$ ii) Cerium ; Due to stable $4f^0$ configuration / Tb ; Due to stable $4f^7$ configuration b) i) Due to ability of oxygen to form multiple bonds to metal ii) HCl is oxidized to chlorine iii) Due to strong interatomic metallic bonding.	½ , ½ ½ , ½ 1 1 1

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