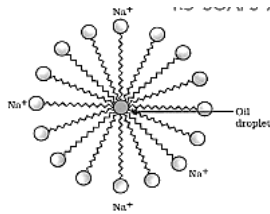
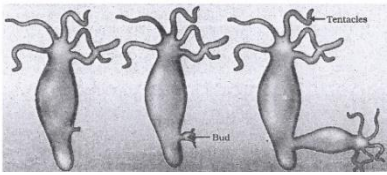


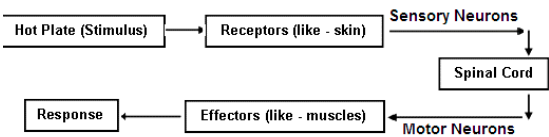
MARKING SCHEME
Secondary School Examination, 2024
SCIENCE (Subject Code–086)
[Paper Code: 31/4/3]

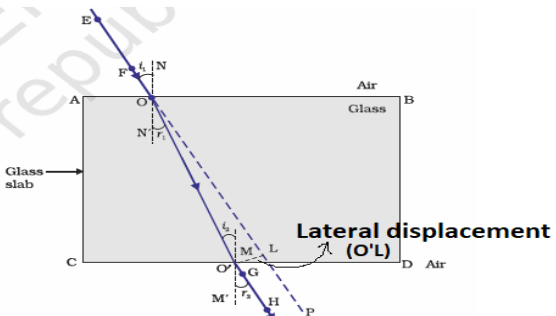
Maximum Marks: 80

Q. No.	EXPECTED ANSWER / VALUE POINTS	Marks	Total Marks
SECTION A			
1	(A)/15000 J	1	1
2	(A)/ HCl, Mg(OH) ₂	1	1
3	(C)/46	1	1
4	(A)/Cotyledon	1	1
5	(B)/Displacement reaction	1	1
6	(D)/Hydrochloric acid and Sulphuric acid	1	1
7	(C)/Copper and Silver	1	1
8	(C)/ Petals only	1	1
9	(B)/ (i)Amino acid, (ii)glucose, (iii)fatty acid and glycerol	1	1
10	(B)/9 and 3	1	1
11	(C)/Abscisic acid	1	1
12	(B)/(ii) and (iii)	1	1
13	(B)/move towards the side AB of the loop	1	1
14	(A)/ $9/4 \times 10^8$ m/s	1	1
15	(C)/Blue	1	1
16	(C)/I and II	1	1
17	(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of Assertion (A).	1	1
18	(A)/ Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of Assertion (A)	1	1
19	(B) /Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of Assertion (A).	1	1
20	(A)/ Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of Assertion (A)	1	1
SECTION B			
21	<p>(a) When water is lost through stomata in the leaves by transpiration, it creates a suction force/transpiration pull. Due to which water is pulled up through xylem of the roots to the leaves.</p> <p style="text-align: center;">OR</p> <p>(b)</p> <ul style="list-style-type: none"> • Heterotrophic /Holozoic Nutrition. • Amoeba takes in food using temporary finger-like projections/pseudopodia of the cell which fuse over the food particle forming a food vacuole. Inside the food vacuole complex substances are broken down into simpler substances. / (award marks if explained diagrammatically) 	<p>2</p> <p>1</p> <p>1</p>	2

22	<ul style="list-style-type: none"> Micelles 	1/2	
		1 1/2	2
23	<p>(a)</p> <p>(i) • HCl gas was evolved</p> <p>(ii) (I) No change in colour (II) Wet blue litmus turns red</p> <ul style="list-style-type: none"> HCl gas is acidic in nature <p style="text-align: center;">OR</p> <p>(b) • $Zn + H_2SO_4 \longrightarrow ZnSO_4 + H_2 (g)$</p> <p style="text-align: right;">(Any other example)</p> <ul style="list-style-type: none"> Hydrogen burns with a pop sound when a burning matchstick is brought near it. 	1/2 1/2 1/2 1/2	
		1 1	2
24	Each parent produces gametes that have half the number of chromosomes. During sexual reproduction, a female gamete fuses with a male gamete to form a zygote. Thus, zygote restores the original number of chromosomes in the progeny ensuring equal contribution of both the parents in the progeny.	2	2
25	<ul style="list-style-type: none"> Black colour As there would not be any particle to scatter light. 	1 1	2
26	$R = \rho \frac{l}{A}$ $\rho = 1.6 \times 10^{-8} \Omega m$ $A = 2 \times (10^{-3} m)^2$ $l = 1000 m$ $\therefore R = (1.6 \times 10^{-8} \Omega m) \times \frac{1000 m}{2 \times (10^{-3} m)^2}$ $= 8.0 \Omega$	1/2 1/2 1	2
SECTION C			
27	<ul style="list-style-type: none"> Ability of the eye lens to adjust its focal length. Image distance remains unchanged Ciliary muscles – <p>While focusing on distant objects ciliary muscles relax, eye lens becomes thin and its focal length increases.</p>	1 1 1/2 1/2	3
28	<p>(a) $CH_3 - COOH + CH_3 - CH_2OH \xrightleftharpoons{Acid} CH_3 - \underset{\underset{O}{ }}{C} - O - CH_2 - CH_3 + H_2O$</p> <p>(b) $CH_3COOC_2H_5 \xrightarrow{NaOH} C_2H_5OH + CH_3COONa$</p> <p>(c) $CH_3 - CH_2OH \xrightarrow[Hot Conc.]{H_2SO_4} CH_2 = CH_2 + H_2O$</p>	1 1 1	3
29	<p>(a) Because a magnetic field exists around the bar magnet</p> <p>(b) Strength of the magnetic field is maximum near the poles of the magnet</p> <p>(c) The lines represent the magnetic field lines</p> <p>(d) Equidistant parallel lines, magnetic field inside the solenoid is uniform</p>	1/2 1 1/2 1/2+1/2	3

30	<p>(a) Two pairs of contrasting characters:</p> <ul style="list-style-type: none"> • Round and wrinkled shape of seed • Violet and white flowers <p>(Any other)</p> <p>(b) (i) No ; Tt (ii) 25% (iii) TT : Tt – 1:2</p>	<p>$\frac{1}{2} + \frac{1}{2}$</p> <p>$\frac{1}{2} ; \frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p>	3				
31	<ul style="list-style-type: none"> • Plaster of Paris; Calcium sulphate hemihydrate • Prepared from gypsum ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$) by heating it at 373K • $\text{CaSO}_4 \cdot \frac{1}{2} \text{H}_2\text{O} + 1\frac{1}{2} \text{H}_2\text{O} \longrightarrow \text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ 	<p>$\frac{1}{2} + \frac{1}{2}$</p> <p>1</p> <p>1</p>	3				
32	<p>(a) In hydra, a bud develops as an outgrowth due to repeated cell division at one specific site. These buds develop into tiny individuals and when fully mature, detach from the parent body and become new independent individuals.</p>  <ul style="list-style-type: none"> • Regenerative cells. <p style="text-align: center;">OR</p> <p>(b)</p> <p>(i) Seminal vesicles and prostate glands:</p> <ul style="list-style-type: none"> • Secrete a fluid for nourishment of sperms. • Secrete a fluid which makes the transport of the sperms easier <p>(ii) Oviduct:</p> <ul style="list-style-type: none"> • Egg is carried from ovary to the womb or uterus. • Site of Fertilization <p>(iii) Testis:</p> <ul style="list-style-type: none"> • Produces sperms • Secretion of hormone – testosterone 	<p>1</p> <p>$1\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2} + \frac{1}{2}$</p> <p>$\frac{1}{2} + \frac{1}{2}$</p> <p>$\frac{1}{2} + \frac{1}{2}$</p>	3				
33	<table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th style="width: 50%;">Biodegradable waste</th> <th style="width: 50%;">Non-biodegradable waste</th> </tr> </thead> <tbody> <tr> <td>Wastes that are broken down by biological processes into simpler substances.</td> <td>Wastes that are not broken down by biological processes into simpler substances.</td> </tr> </tbody> </table> <ul style="list-style-type: none"> • Harmful effects: <ul style="list-style-type: none"> - Excessive use cause pollution. - Pesticides enter the food chain and cause biomagnification in humans and other animals. - Clogging of drains. - Death of cattle due to ingestion of plastics <p style="text-align: right;">(any two)</p>	Biodegradable waste	Non-biodegradable waste	Wastes that are broken down by biological processes into simpler substances.	Wastes that are not broken down by biological processes into simpler substances.	<p>2</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p>	3
Biodegradable waste	Non-biodegradable waste						
Wastes that are broken down by biological processes into simpler substances.	Wastes that are not broken down by biological processes into simpler substances.						
SECTION D							
34	<p>(a)</p> <p>(i) • The pathway in which impulses travel during the reflex action is called a reflex arc.</p> <ul style="list-style-type: none"> • Because the thinking part of the brain is not fast enough/for quick 	<p>1</p> <p>$\frac{1}{2}$</p>					

	<p>response to avoid injury.</p> <ul style="list-style-type: none"> Reflex arc :  <p>(ii) Peripheral Nervous System Components : Cranial Nerves; Spinal Nerves</p> <p style="text-align: center;">OR</p> <p>(b)</p> <p>(i) •Touch</p> <ul style="list-style-type: none"> The shape of the leaves changes by changing the amount of water in them. No <p>(ii) Growth of a part of plant in response to the pull of earth or gravity is called geotropism.</p> <ul style="list-style-type: none"> Positive geotropism – Movement of plant part towards the earth gravity. Example – Roots grow downwards Negative geotropism – Movement of plant part away from the force of gravity. Example – Shoots grow upwards. 	<p>1+½</p> <p>1 ½; ½</p> <p>½ 1 ½</p> <p>1 ½+½ ½+½</p>	<p>5</p>
<p>35</p>	<p>(a) A chemical reaction involves the breaking and making of bonds between atoms to produce new substances. / when reactant changes to products.</p> <p>(i) Add lead nitrate solution to potassium iodide solution taken in a test tube. The colour changes from colourless solution to yellow ppt. /</p> $\text{Pb (NO}_3)_2 + 2\text{KI} \longrightarrow \text{PbI}_2 \downarrow + 2\text{KNO}_3$ <p style="text-align: center; color: orange;">Yellow</p> <p style="text-align: right;">(or any example)</p> <p>(ii) Calcium oxide reacts vigorously with water to produce slaked lime (calcium hydroxide) releasing a large amount of heat. /</p> $\text{CaO(s)} + \text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{Heat}$ <p style="text-align: center;">(Quick lime) (Slaked lime)</p> <p style="text-align: right;">(or any example)</p> <p>(Deduct ½ marks if change in colour or heat is not mentioned in the reaction)</p> <p style="text-align: center;">OR</p> <p>(b)</p> <p>(i) •A reactant breaks down to give two or more products. /A reaction which requires energy to split a compound or reactant in two or more simple substances.</p> <p>(I) Water splits into hydrogen gas and oxygen gas.</p> <ul style="list-style-type: none"> Electrical energy <p>(II) Silver bromide decomposes into silver and bromine</p> <ul style="list-style-type: none"> Light energy <p>(ii)</p> <p>(I) Formation of calcium oxide:</p> $\text{CaCO}_3\text{(s)} \xrightarrow{\text{Heat}} \text{CaO(s)} + \text{CO}_2\text{(g)}$ <ul style="list-style-type: none"> It is an endothermic reaction/decomposition reaction. 	<p>1</p> <p>1+1</p> <p>1+1</p> <p>1</p> <p>½ ½</p> <p>½ ½</p> <p>½ ½</p>	

	<p>(II) Formation of calcium hydroxide:</p> $CaO + H_2O \rightarrow Ca(OH)_2 + \text{Heat}$ <ul style="list-style-type: none"> It is exothermic/combination reaction 	$\frac{1}{2}$ $\frac{1}{2}$	5
36	<p>(a) (i)</p> <ul style="list-style-type: none"> The angle of incidence is equal to the angle of reflection. The incident ray, the normal to the mirror at the point of incidence and the reflected ray, all lie in the same plane. <p>(ii) $u = -15 \text{ cm}$, $f = -10 \text{ cm}$ (concave mirror) $h = 5.0 \text{ cm}$</p> <p>Mirror formula $\frac{1}{f} = \frac{1}{v} + \frac{1}{u}$</p> $\frac{1}{v} = \frac{-1}{10 \text{ cm}} + \frac{1}{15 \text{ cm}} = \frac{-1}{30 \text{ cm}}$ <p>or $v = -30 \text{ cm}$. The screen must be placed at a distance of 30 cm from the mirror in front of it</p> $(m) = \frac{h'}{h} = -\frac{v}{u}$ $h = \frac{-v}{u} \times h = -\frac{-30 \text{ cm}}{-15 \text{ cm}} \times 5 \text{ cm} = -10 \text{ cm}$ <p style="text-align: center;">OR</p> <p>(b)(i)</p> <ul style="list-style-type: none"> The incident ray, the refracted ray and the normal to the interface of two transparent media at the point of incidence, all lie in the same plane. The ratio of sine of angle of incidence to the sine of angle of refraction is a constant, for the light of a given colour and for the given pair of media. / $\frac{\sin i}{\sin r} = \text{constant}$ <p>(ii)</p>  <ul style="list-style-type: none"> The emergent ray is parallel to the incident ray. Labelling of lateral displacement <p style="text-align: center;">(If labelling is not done deduct 1/2 marks)</p>	<p>1+1</p> $\frac{1}{2}$	$\frac{1}{2}$
	SECTION E		
37	<p>(a) $R_s = 4 \Omega + 6 \Omega + 16 \Omega = 26 \Omega$</p> <p>(b) $\frac{1}{R_p} = \frac{1}{8 \Omega} + \frac{1}{8 \Omega} = \frac{1}{4} \Omega$ $R_p = 4 \Omega$</p> <p>(c) (i) Total resistance = $26 \Omega + 4 \Omega = 30 \Omega$ Potential difference = $V = 6V$ Current $I = \frac{V}{R}$ $\frac{6}{30} = \frac{1}{5} \text{ A}$ or 0.2 A.</p>	<p>1</p> <p>1</p> <p>1</p> $\frac{1}{2}$ $\frac{1}{2}$	

	OR		
	<p>(c)(ii) 16 Ω Justification: According to Ohm's law when same current flows, the potential difference across a higher resistance is always higher./ Potential difference across 16 $\Omega = V = IR = 0.2 \times 16 = 3.2V$ Potential difference across 8 $\Omega = V = IR_{(total)} = 0.2 \times 4 = 0.8V$</p>	1 1	4
38	<p>(a) In the test tube containing magnesium. (b) All three metals react with HCl because they are more reactive than hydrogen. (Award marks if student write any less reactive metal with reason) (c) (i) Because HNO_3 is a strong oxidizing agent and oxidizes the H_2 produced to water. <ul style="list-style-type: none"> • Ultimate products are water, oxides of nitrogen. <p style="text-align: center;">OR</p> (c) (ii) • Displacement Reaction • If metal X displaces metal Y from its salt solution it is more reactive than Y or vice versa.</p>	1 1 1 1 1 1	4
39	<p>(a) (i) Renal Artery (ii) Glomerulus</p> <p>(b) • Urinary bladder • Nervous control</p> <p>(c) (i) Filtration: Nitrogenous wastes such as urea or uric acid are removed Reabsorption: Glucose, amino acids, salts/some useful materials and major amounts of water reabsorbed</p> <p style="text-align: center;">OR</p> (c) (ii) Tubular part of nephron. <ul style="list-style-type: none"> • The amount of water absorbed depends on : -how much water is there in the body. -how much dissolved waste is there to be excreted. 	1/2 1/2 1/2 1/2 1/2+1/2 1/2+1/2 1 1/2 1/2	4