

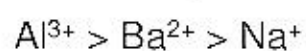
CHEMISTRY

SECTION-A

51. Given below are two statements

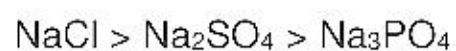
Statement I:

In the coagulation of a negative sol, the flocculating power of the three given ions is in the order



Statement II:

In the coagulation of a positive sol, the flocculating power of the three given salts is in the order



In the light of the above statements, choose the most **appropriate** answer from the options given below

- (1) Statement I is correct but Statement II is incorrect.
- (2) Statement I is incorrect but Statement II is correct.
- (3) Both Statement I and Statement II are correct.
- (4) Both Statement I and Statement II are incorrect.

Answer (1)

52. Which one is not correct mathematical equation for Dalton's Law of partial pressure? Here p = total pressure of gaseous mixture

(1) $p_i = \chi_i p$,

where p_i = partial pressure of i^{th} gas

χ_i = mole fraction of i^{th} gas in gaseous mixture

(2) $p_i = \chi_i p_i^{\circ}$,

where χ_i = mole fraction of i^{th} gas in gaseous mixture

p_i° = pressure of i^{th} gas in pure state

(3) $p = p_1 + p_2 + p_3$

(4) $p = n_1 \frac{RT}{V} + n_2 \frac{RT}{V} + n_3 \frac{RT}{V}$

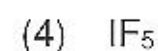
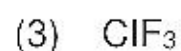
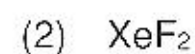
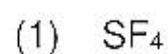
Answer (2)

53. Which statement regarding polymers is **not correct**?

- (1) Thermoplastic polymers are capable of repeatedly softening and hardening on heating and cooling respectively
- (2) Thermosetting polymers are reusable
- (3) Elastomers have polymer chains held together by weak intermolecular forces
- (4) Fibers possess high tensile strength

Answer (2)

54. Amongst the following which one will have maximum 'lone pair - lone pair' electron repulsions?



Answer (2)

55. Identify the incorrect statement from the following
- (1) Ionisation enthalpy of alkali metals decreases from top to bottom in the group.
 - (2) Lithium is the strongest reducing agent among the alkali metals.
 - (3) Alkali metals react with water to form their hydroxides.
 - (4) The oxidation number of K in KO_2 is +4.

Answer (4)

56. Given below are two statements

Statement I

The boiling points of the following hydrides of group 16 elements increases in the order –
 $\text{H}_2\text{O} < \text{H}_2\text{S} < \text{H}_2\text{Se} < \text{H}_2\text{Te}$

Statement II

The boiling points of these hydrides increase with increase in molar mass.

In the light of the above statements, choose the most appropriate answer from the options given below :

- (1) **Statement I** is correct but **Statement II** is incorrect
- (2) **Statement I** is incorrect but **Statement II** is correct
- (3) Both **Statement I** and **Statement II** are correct
- (4) Both **Statement I** and **Statement II** are incorrect

Answer (4)

57. The IUPAC name of an element with atomic number 119 is

- | | |
|----------------|-----------------|
| (1) unununnium | (2) ununoctium |
| (3) ununennium | (4) unnilennium |

Answer (3)

58. Given below are two statements

Statement I:

The acidic strength of monosubstituted nitrophenol is higher than phenol because of electron withdrawing nitro group.

Statement II:

o-nitrophenol, *m*-nitrophenol and *p*-nitrophenol will have same acidic strength as they have one nitro group attached to the phenolic ring.

In the light of the above statements, choose the **most appropriate** answer from the options given below:

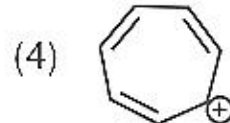
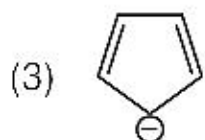
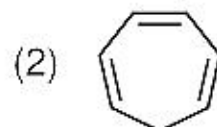
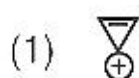
- (1) **Statement I** is correct but **Statement II** is incorrect.
- (2) **Statement I** is incorrect but **Statement II** is correct.
- (3) Both **Statement I** and **Statement II** are correct.
- (4) Both **Statement I** and **Statement II** are incorrect.

59. Which amongst the following is **incorrect** statement?

- (1) H_2^+ ion has one electron
- (2) O_2^- ion is diamagnetic
- (3) The bond orders of O_2^+ , O_2 , O_2^- and O_2^{2-} are 2.5, 2, 1.5 and 1, respectively
- (4) C_2 molecule has four electrons in its two degenerate π molecular orbitals

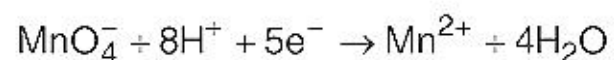
Answer (2)

60. Which compound amongst the following is **not** an aromatic compound?

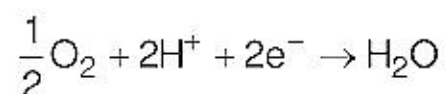


Answer (2)

61. Given below are half cell reactions:



$$E_{\text{Mn}^{2+}/\text{MnO}_4^-}^{\circ} = -1.510 \text{ V}$$



$$E_{\text{O}_2/\text{H}_2\text{O}}^{\circ} = +1.223 \text{ V}$$

Will the permanganate ion, MnO_4^- liberate O_2 from water in the presence of an acid?

(1) Yes, because $E_{\text{cell}}^{\circ} = + 2.733 \text{ V}$

(2) No, because $E_{\text{cell}}^{\circ} = - 2.733 \text{ V}$

(3) Yes, because $E_{\text{cell}}^{\circ} = + 0.287 \text{ V}$

(4) No, because $E_{\text{cell}}^{\circ} = - 0.287 \text{ V}$

Answer (3)

62. The IUPAC name of the complex-



(1) dicyanosilver(I) diaquaargentate(I)

(2) diaquasilver(I) dicyanidoargentate(I)

(3) dicyanosilver(II) diaquaargentate(II)

(4) diaquasilver(II) dicyanidoargentate(II)

Answer (2)

63. Given below are two statements

Statement I:

Primary aliphatic amines react with HNO_2 to give unstable diazonium salts.

Statement II:

Primary aromatic amines react with HNO_2 to form diazonium salts which are stable even above 300 K. In the light of the above statements, choose the **most appropriate** answer from the options given below

(1) Statement I is correct but Statement II is incorrect.

(2) Statement I is incorrect but Statement II is correct.

(3) Both Statement I and Statement II are correct.

(4) Both Statement I and Statement II are incorrect.

Answer (1)

64. Which of the following statement is not correct about diborane?

(1) The four terminal Hydrogen atoms and the two Boron atoms lie in one plane.

(2) Both the Boron atoms are sp^2 hybridised.

(3) There are two 3-centre-2-electron bonds.

(4) The four terminal B-H bonds are two centre two electron bonds.

Answer (2)

65. Match List-I with List-II.

List – I (Products formed)	List – II (Reaction of carbonyl compound with)
(a) Cyanohydrin	(i) NH_2OH
(b) Acetal	(ii) RNH_2
(c) Schiff's base	(iii) alcohol
(d) Oxime	(iv) HCN

Choose the correct answer from the options given below

- (1) (a) – (i), (b) – (iii), (c) – (ii), (d) – (iv) (2) (a) – (iv), (b) – (iii), (c) – (ii), (d) – (i)
(3) (a) – (iii), (b) – (iv), (c) – (ii), (d) – (i) (4) (a) – (ii), (b) – (iii), (c) – (iv), (d) – (i)

Answer (2)

66. Choose the correct statement:

- (1) Diamond is sp^3 hybridised and graphite is sp^2 hybridized.
(2) Both diamond and graphite are used as dry lubricants.
(3) Diamond and graphite have two dimensional network.
(4) Diamond is covalent and graphite is ionic.

Answer (1)

67. The pH of the solution containing 50 mL each of 0.10 M sodium acetate and 0.01 M acetic acid is

[Given pK_a of $\text{CH}_3\text{COOH} = 4.57$]

- (1) 4.57 (2) 2.57
(3) 5.57 (4) 3.57

Answer (3)

68. Gadolinium has a low value of third ionisation enthalpy because of

- (1) high electronegativity (2) high basic character
(3) small size (4) high exchange enthalpy

Answer (4)

69. In one molal solution that contains 0.5 mole of a solute, there is

- (1) 100 mL of solvent (2) 1000 g of solvent
(3) 500 mL of solvent (4) 500 g of solvent

Answer (4)

70. Given below are two statements :

Statement I : The boiling points of aldehydes and ketones are higher than hydrocarbons of comparable molecular masses because of weak molecular association in aldehydes and ketones due to dipole - dipole interactions.

Statement II : The boiling points of aldehydes and ketones are lower than the alcohols of similar molecular masses due to the absence of H-bonding.

In the light of the above statements, choose the **most appropriate** answer from the given below

- (1) **Statement I** is correct but **Statement II** is incorrect
 (2) **Statement I** is incorrect but **Statement II** is correct
 (3) Both **Statement I** and **Statement II** are correct
 (4) Both **Statement I** and **Statement II** are incorrect

Answer (3)

71. At 298 K, the standard electrode potentials of $\text{Cu}^{2+} / \text{Cu}$, $\text{Zn}^{2+} / \text{Zn}$, $\text{Fe}^{2+} / \text{Fe}$ and Ag^+ / Ag are 0.34 V, -0.76 V, -0.44 V and 0.80 V, respectively.

On the basis of standard electrode potential, predict which of the following reaction cannot occur?

- (1) $\text{FeSO}_4(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Fe}(\text{s})$ (2) $2\text{CuSO}_4(\text{aq}) + 2\text{Ag}(\text{s}) \rightarrow 2\text{Cu}(\text{s}) + \text{Ag}_2\text{SO}_4(\text{aq})$
 (3) $\text{CuSO}_4(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu}(\text{s})$ (4) $\text{CuSO}_4(\text{aq}) + \text{Fe}(\text{s}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu}(\text{s})$

Answer (2)

72. Match **List-I** with **List-II**

List-I

- (a) Li
 (b) Na
 (c) KOH
 (d) Cs

List-II

- (i) absorbent for carbon dioxide
 (ii) electrochemical cells
 (iii) coolant in fast breeder reactors
 (iv) photoelectric cell

Choose the correct answer from the options given below :

- (1) (a) - (i), (b) - (iii), (c) - (iv), (d) - (ii) (2) (a) - (ii), (b) - (iii), (c) - (i), (d) - (iv)
 (3) (a) - (iv), (b) - (i), (c) - (iii), (d) - (ii) (4) (a) - (iii), (b) - (iv), (c) - (ii), (d) - (i)

Answer (2)

73. Match **List-I** with **List-II**.

List-I

(Drug class)

- (a) Antacids
 (b) Antihistamines
 (c) Analgesics
 (d) Antimicrobials

List-II

(Drug molecule)

- (i) Salvarsan
 (ii) Morphine
 (iii) Cimetidine
 (iv) Seldane

Choose the correct answer from the options given below :

- (1) (a) - (i), (b) - (iv), (c) - (ii), (d) - (iii) (2) (a) - (iv), (b) - (iii), (c) - (i), (d) - (ii)
 (3) (a) - (iii), (b) - (ii), (c) - (iv), (d) - (i) (4) (a) - (iii), (b) - (iv), (c) - (ii), (d) - (i)

Answer (4)

74. $\text{RMgX} + \text{CO}_2 \xrightarrow[\text{ether}]{\text{dry}} \text{Y} \xrightarrow{\text{H}_3\text{O}^+} \text{RCOOH}$

What is Y in the above reaction?

- (1) $\text{RCOO}^- \text{X}^+$ (2) $(\text{RCOO})_2\text{Mg}$
 (3) $\text{RCOO}^- \text{Mg}^+ \text{X}$ (4) $\text{R}_3\text{CO}^- \text{Mg}^+ \text{X}$

Answer (3)

75. Given below are two statements: one is labelled as **Assertion (A)** and the other is labelled as **Reason (R)**.

Assertion (A): ICl is more reactive than I₂.

Reason (R): I-Cl bond is weaker than I-I bond.

In the light of the above statements, choose the most **appropriate** answer from the options given below:

- (1) (A) is correct but (R) is not correct
- (2) (A) is not correct but (R) is correct
- (3) Both (A) and (R) are correct and (R) is the correct explanation of (A).
- (4) Both (A) and (R) are correct but (R) is not the correct explanation of (A).

Answer (3)

76. The **incorrect** statement regarding chirality is

- (1) Enantiomers are superimposable mirror images on each other
- (2) A racemic mixture shows zero optical rotation
- (3) S_N1 reaction yields 1 : 1 mixture of both enantiomers
- (4) The product obtained by S_N2 reaction of haloalkane having chirality at the reactive site shows inversion of configuration

Answer (1)

77. Match List-I with List-II.

List – I

(Hydrides)

- (a) MgH₂
- (b) GeH₄
- (c) B₂H₆
- (d) HF

List – II

(Nature)

- (i) Electron precise
- (ii) Electron deficient
- (iii) Electron rich
- (iv) Ionic

Choose the correct answer from the options given below

- (1) (a) – (i), (b) – (ii), (c) – (iv), (d) – (iii)
- (2) (a) – (ii), (b) – (iii), (c) – (iv), (d) – (i)
- (3) (a) – (iv), (b) – (i), (c) – (ii), (d) – (iii)
- (4) (a) – (iii), (b) – (i), (c) – (ii), (d) – (iv)

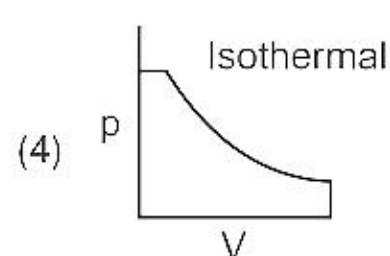
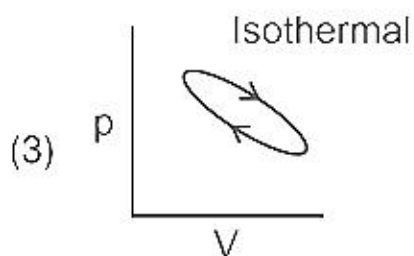
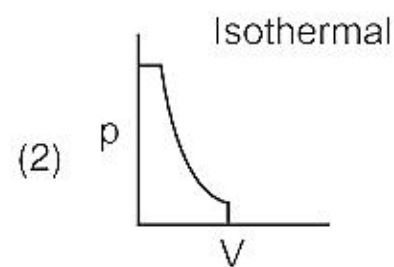
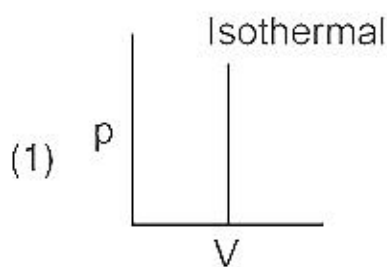
Answer (3)

78. The **incorrect** statement regarding enzymes is

- (1) Enzymes are polysaccharides.
- (2) Enzymes are very specific for a particular reaction and substrate.
- (3) Enzymes are biocatalysts.
- (4) Like chemical catalysts enzymes reduce the activation energy of bio processes.

Answer (1)

79. Which of the following p-V curve represents maximum work done?



Answer (4)

80. Given below are two statements: one is labelled as **Assertion (A)** and the other is labelled as **Reason (R)**.

Assertion (A):

In a particular point defect, an ionic solid is electrically neutral, even if few of its cations are missing from its unit cells.

Reason (R):

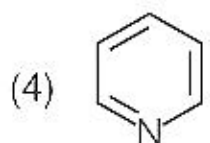
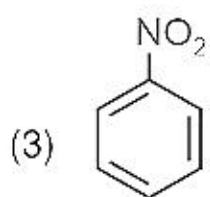
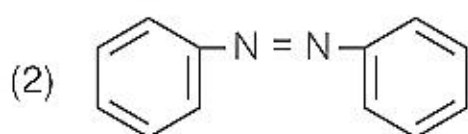
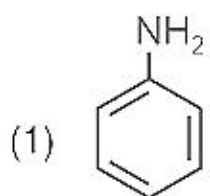
In an ionic solid, Frenkel defect arises due to dislocation of cation from its lattice site to interstitial site, maintaining overall electrical neutrality.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) (A) is correct but (R) is not correct
- (2) (A) is not correct but (R) is correct
- (3) Both (A) and (R) are correct and (R) is the correct explanation of (A)
- (4) Both (A) and (R) are correct but (R) is not the correct explanation of (A)

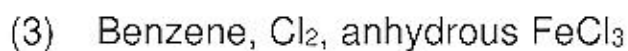
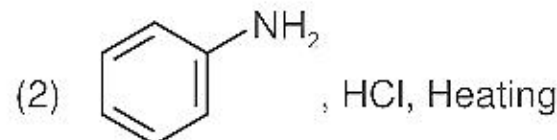
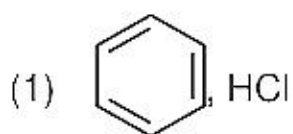
Answer (4)

81. The Kjeldahl's method for the estimation of nitrogen can be used to estimate the amount of nitrogen in which one of the following compounds?



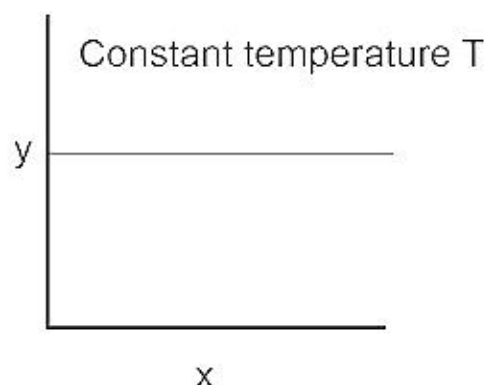
Answer (1)

82. Which of the following sequence of reactions is suitable to synthesize chlorobenzene?



Answer (3)

83. The given graph is a representation of kinetics of a reaction.



The y and x axes for zero and first order reactions, respectively are

- (1) zero order (y = rate and x = concentration), first order (y = t_{1/2} and x = concentration)
- (2) zero order (y = rate and x = concentration), first order (y = rate and x = t_{1/2})
- (3) zero order (y = concentration and x = time), first order (y = t_{1/2} and x = concentration)
- (4) zero order (y = concentration and x = time), first order (y = rate constant and x = concentration)

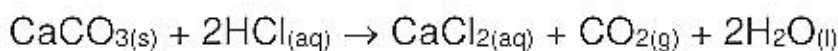
Answer (1)

84. Identify the **incorrect** statement from the following.

- (1) In an atom, all the five 3d orbitals are equal in energy in free state.
- (2) The shapes of d_{xy}, d_{yz} and d_{zx} orbitals are similar to each other; and d_{x²-y²} and d_{z²} are similar to each other.
- (3) All the five 5d orbitals are different in size when compared to the respective 4d orbitals.
- (4) All the five 4d orbitals have shapes similar to the respective 3d orbitals.

Answer (2)

85. What mass of 95% pure CaCO₃ will be required to neutralise 50 mL of 0.5 M HCl solution according to the following reaction?

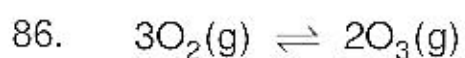


[Calculate upto second place of decimal point]

- (1) 3.65 g
- (2) 9.50 g
- (3) 1.25 g
- (4) 1.32 g

Answer (4)

SECTION-B



for the above reaction at 298 K, K_c is found to be 3.0×10^{-59} . If the concentration of O_2 at equilibrium is 0.040 M then concentration of O_3 in M is

- (1) 2.4×10^{31} (2) 1.2×10^{21}
 (3) 4.38×10^{-32} (4) 1.9×10^{-63}

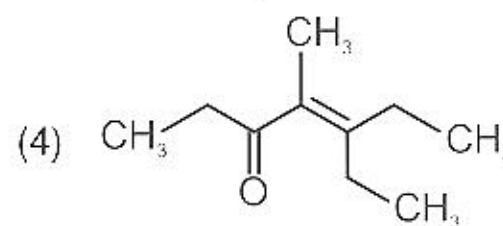
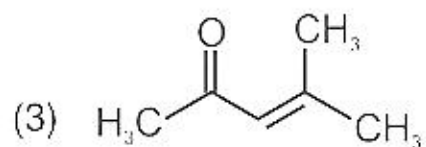
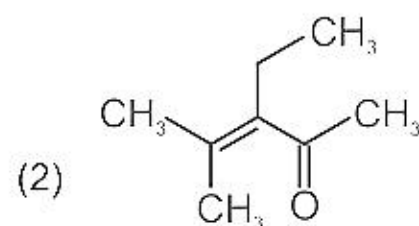
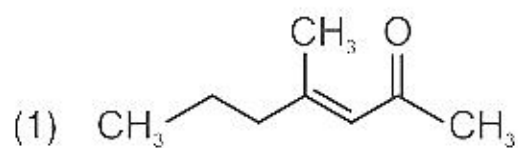
Answer (3)

87. Compound X on reaction with O_3 followed by $\text{Zn}/\text{H}_2\text{O}$ gives formaldehyde and 2-methyl propanal as products. The compound X is

- (1) 2-Methylbut-2-ene (2) Pent-2-ene
 (3) 3-Methylbut-1-ene (4) 2-Methylbut-1-ene

Answer (3)

88. Which one of the following is not formed when acetone reacts with 2-pentanone in the presence of dilute NaOH followed by heating?



Answer (4)

89. For a first order reaction $\text{A} \rightarrow \text{Products}$, initial concentration of A is 0.1 M, which becomes 0.001 M after 5 minutes. Rate constant for the reaction in min^{-1} is

- (1) 0.4606 (2) 0.2303
 (3) 1.3818 (4) 0.9212

Answer (4)

90. Given below are two statements:

Statement I:

In Lucas test, primary, secondary and tertiary alcohols are distinguished on the basis of their reactivity with conc. $\text{HCl} + \text{ZnCl}_2$, known as Lucas Reagent.

Statement II:

Primary alcohols are most reactive and immediately produce turbidity at room temperature on reaction with Lucas Reagent.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Statement I is correct but Statement II is incorrect

- (2) Statement I is incorrect but Statement II is correct
 (3) Both Statement I and Statement II are correct
 (4) Both Statement I and Statement II are incorrect

Answer (1)

91. Match List-I with List-II.

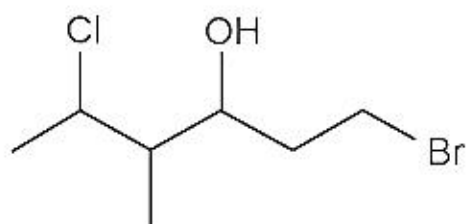
List-I (Ores)	List-II (Composition)
(a) Haematite	(i) Fe_3O_4
(b) Magnetite	(ii) ZnCO_3
(c) Calamine	(iii) Fe_2O_3
(d) Kaolinite	(iv) $[\text{Al}_2(\text{OH})_4\text{Si}_2\text{O}_5]$

Choose the correct answer from the options given below:

- (1) (a)-(iii), (b)-(i), (c)-(iv), (d)-(ii) (2) (a)-(i), (b)-(iii), (c)-(ii), (d)-(iv)
 (3) (a)-(i), (b)-(ii), (c)-(iii), (d)-(iv) (4) (a)-(iii), (b)-(i), (c)-(ii), (d)-(iv)

Answer (4)

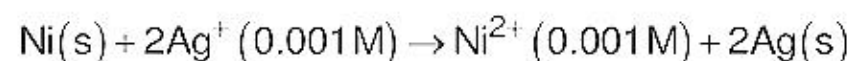
92. The correct IUPAC name of the following compound is



- (1) 1-bromo-4-methyl-5-chlorohexan-3-ol (2) 6-bromo-4-methyl-2-chlorohexan-4-ol
 (3) 1-bromo-5-chloro-4-methylhexan-3-ol (4) 6-bromo-2-chloro-4-methylhexan-4-ol

Answer (3)

93. Find the emf of the cell in which the following reaction takes place at 298 K



(Given that $E_{\text{cell}}^\circ = 10.5\text{ V}$, $\frac{2.303 RT}{F} = 0.059$ at 298 K)

- (1) 0.9615 V (2) 1.05 V
 (3) 1.0385 V (4) 1.385 V

Answer (NA)

94. The order of energy absorbed which is responsible for the color of complexes

- (A) $[\text{Ni}(\text{H}_2\text{O})_2(\text{en})_2]^{2+}$
 (B) $[\text{Ni}(\text{H}_2\text{O})_4(\text{en})]^{2+}$ and
 (C) $[\text{Ni}(\text{en})_3]^{2+}$

is

- (1) (C) > (A) > (B) (2) (B) > (A) > (C)
 (3) (A) > (B) > (C) (4) (C) > (B) > (A)

Answer (1)

95. The pollution due to oxides of sulphur gets enhanced due to the presence of:

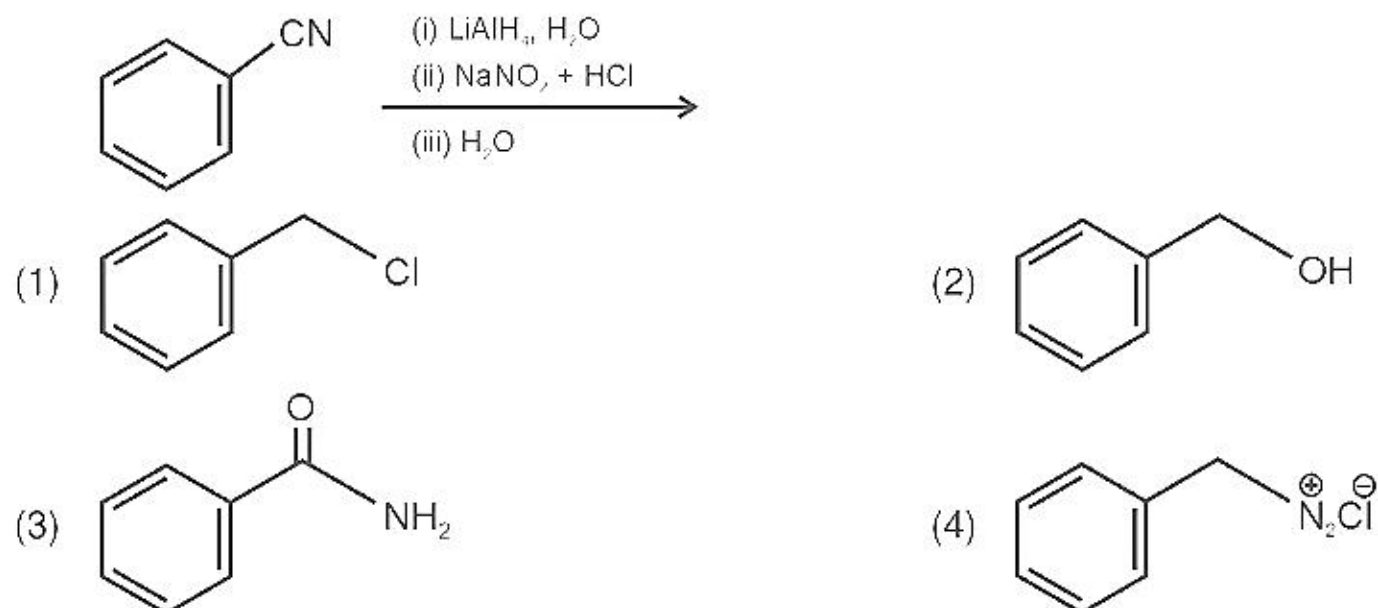
- (a) particulate matter (b) ozone
(c) hydrocarbons (d) hydrogen peroxide

Choose the most appropriate answer from the options given below:

- (1) (b), (c), (d) only (2) (a), (c), (d) only
(3) (a), (d) only (4) (a), (b), (d) only

Answer (4)

96. The product formed from the following reaction sequence is



Answer (2)

97. If radius of second Bohr orbit of the He^+ ion is 105.8 pm, what is the radius of third Bohr orbit of Li^{2+} ion?

- (1) 1.587 pm (2) 158.7 Å
(3) 158.7 pm (4) 15.87 pm

Answer (3)

98. In the neutral or faintly alkaline medium, KMnO_4 oxidises iodide into iodate. The change in oxidation state of manganese in this reaction is from

- (1) +7 to +3 (2) +6 to +5
(3) +7 to +4 (4) +6 to +4

Answer (3)

99. Copper crystallises in fcc unit cell with cell edge length of 3.608×10^{-8} cm. The density of copper is 8.92 g cm^{-3} . Calculate the atomic mass of copper.

- (1) 60 u (2) 65 u
(3) 63.1 u (4) 31.55 u

Answer (3)

100. A 10.0 L flask contains 64 g of oxygen at 27°C . (Assume O_2 gas is behaving ideally). The pressure inside the flask in bar is (Given $R = 0.0831 \text{ L bar K}^{-1} \text{ mol}^{-1}$)

- (1) 49.8 (2) 4.9
(3) 2.5 (4) 498.6

Answer (2)