



JEE (MAIN) 2024

MEMORY BASED QUESTIONS & SOLUTIONS

SHIFT-2

DATE & DAY: 31st January 2024 & Wednesday

PAPER-1

Duration: 3 Hrs.

Time: 03:00 PM - 06:00 PM

SUBJECT: CHEMISTRY

ADMISSIONS OPEN FOR CLASS 12+

ACADEMIC SESSION 2024-25



TARGET: JEE (ADV.) 2024

For Class XII Passed Student
VISHESH COURSE
MODE: OFFLINE/ONLINE

CLASS STARTS
08th APRIL, 2024



TARGET: JEE (MAIN) 2024

For Class XII Passed Student
ABHYAAS COURSE
MODE: OFFLINE/ONLINE

CLASS STARTS
08th APRIL, 2024

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COURSE COMMENCEMENT: 5th FEBRUARY 2024

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SCAN TO
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| JEE(Main) 2024 | DATE : 31-01-2024 (SHIFT-2) | PAPER-1 | MEMORY BASED | CHEMISTRY

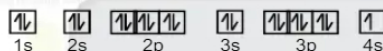
PART : CHEMISTRY

1. Find the correct set of quantum number for the last electron of potassium.

- (1) $n = 4$ $\ell = 2$ $m = +2$ $s = +\frac{1}{2}$ (2) $n = 2$ $\ell = 0$ $m = 0$ $s = +\frac{1}{2}$
 (3) $n = 3$ $\ell = 0$ $m = 0$ $s = +\frac{1}{2}$ (4) $n = 4$ $\ell = 0$ $m = 0$ $s = +\frac{1}{2}$

Ans. (4)

Sol. K = 19



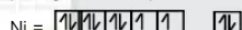
$$\rightarrow n = 4, \ell = 0, m = 0, s = +\frac{1}{2}$$

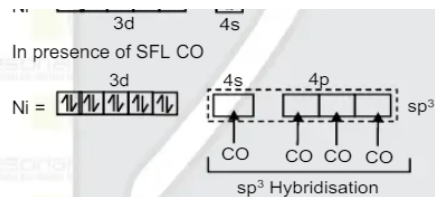
2. Which of the following is correct ?

- (1) $[\text{NiCl}_4]^{2-}$ diamagnetic, $[\text{Ni}(\text{CO})_4]$ diamagnetic
 (2) $[\text{Ni}(\text{CO})_4]$ diamagnetic, $[\text{Ni}(\text{Cl})_4]^{2-}$ Paramagnetic
 (3) $[\text{NiCl}_4]^{2-}$ paramagnetic, $[\text{Ni}(\text{CO})_4]$ Paramagnetic
 (4) $[\text{NiCl}_4]^{2-}$ diamagnetic $[\text{Ni}(\text{CO})_4]$ Paramagnetic

Ans. (2)

Sol. $\text{Ni}(\text{CO})_4 \rightarrow \text{Ni}(0) \rightarrow$ In Presence of strong field ligand CO \rightarrow more interaction and Δ value high so pairing will takes place.

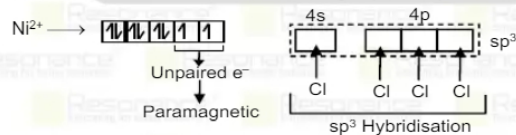




No unpaired electron - So diamagnetic compound.

But in $[NiCl_4]^{2-} \rightarrow Ni^{+2}$

$Cl^- \rightarrow WFL \rightarrow \Delta_0$ Value low \rightarrow high spin complex



So Option (2) is Correct Answer.

3. Which of the following is least ionic ?

(1) $BaCl_2$ (2) KCl (3) $AgCl$ (4) $CoCl_2$

Ans. (3)

Sol. Order of ionic character = $BaCl_2 > KCl > CoCl_2 > AgCl$.

Ag^+ due to pseudo inert gas configuration have high polarizing power.

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4. Statement-I : 13th group element hydrolyse due to covalent nature.

Statement-II : On hydrolysis Al give $[Al(H_2O)_6]^{3+}$

- (1) Both statement I and II are correct
 (2) Statement I is correct and statement II is incorrect
 (3) Statement I is incorrect and statement II is correct
 (4) Both statement I and II are incorrect

Ans. (1)

Sol. 13th group element hydrolyse due to covalent character and on hydrolysis Al give $[Al(H_2O)_6]^{3+}$.

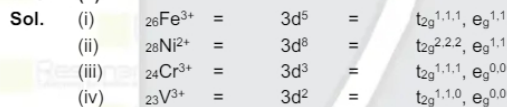
5.

| | List-I | | List-II |
|-------|---------------------|-----|-----------------------------|
| | Complex | | Electronic configuration |
| (i) | $[Fe(H_2O)_6]^{3+}$ | (P) | $t_{2g}^{2,2,2}, e_g^{1,1}$ |
| (ii) | $[Ni(H_2O)_6]^{2+}$ | (Q) | $t_{2g}^{1,1,1}, e_g^{1,1}$ |
| (iii) | $[Cr(H_2O)_6]^{3+}$ | (R) | $t_{2g}^{1,1,0}, e_g^{0,0}$ |
| (iv) | $[V(H_2O)_6]^{3+}$ | (S) | $t_{2g}^{1,1,1}, e_g^{0,0}$ |

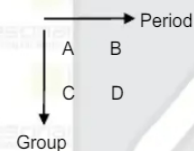
Identify correct match:

- (1) i-Q, ii-P, iii-S, iv-R (2) i-P, ii-Q, iii-R, iv-S (3) i-Q, ii-P, iii-R, iv-S (4) i-P, ii-R, iii-Q, iv-S

Ans. (1)



6.



Where A, B, C and D are elements in periodic table

Which of the following order is correct ?

- (1) Atomic radius: $A < B < C < D$ (2) Metallic radius: $A < B < D < C$
 (3) Ionic radius: $B^+ < A^+ < D^+ < C^+$ (4) None of these.

Ans. (3)

Sol. (i) On moving left to right atomic radius decrease. So atomic radius order is $B < A$.

(ii) Metallic radius is also decrease on moving left to right so order $\Rightarrow B < A$.

(iii) Order of ionic radius $B^+ < A^+ < D^+ < C^+$.

7. (A) Mn_2O_7 is an oil at room temperature (B) V_2O_5 reacts with acid to give VO_2^+
 (C) CrO is a basic oxide (D) V_2O_5 does not react with acids
 Choose the correct answer
 (1) A, B and C (2) B, C and D only (3) A only (4) B and C only

Ans. (1)

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Sol. Ref. NCERT 4.4.1

- Mn_2O_7 is a covalent green oil. So (A) is correct.
- V_2O_5 is amphoteric, though mainly acidic, it gives VO_4^{3-} as well as VO_2^+ salts. So (B) is correct.
- CrO is basic because Cr in CrO is +2 oxidation state. So Cr in +2 oxidation state shows basic character. So (C) is correct.

8. **Statement-I** : In 15th group hydrides reducing character decreases from NH_3 to BiH_3 .

Statement-II : E_2O_5 is more acidic than E_2O_3 (where E is the 15th group elements)

- (1) Statement I is incorrect and statement II is correct
 (2) Statement I is correct and statement II is incorrect
 (3) Both statement I and II are correct
 (4) Both statement I and II are incorrect

Ans. (1)

- Sol. (1) Reducing character of 15th group hydrides increases from NH_3 to BiH_3 .
 (2) The oxide in the higher oxidation state of element is more acidic than that of lower oxidation state.

9. **Statement-I** : In the reduction of permanganate ion to manganate ion, one e^- is involved.

Statement-II : $CrO_4^{2-} \xrightarrow{H^+}$ Product

In product Oxidation number of Cr is 6.

- (1) Statement I is incorrect and statement II is correct
 (2) Statement I is correct and statement II is incorrect
 (3) Both statement I and II are correct
 (4) Both statement I and II are incorrect

Ans. (1)

- Sol. $MnO_4^- + e^- \longrightarrow MnO_4^{2-}$
 $2CrO_4^{2-} + 2H^+ \longrightarrow Cr_2O_7^{2-} + H_2O$
 In product $Cr_2O_7^{2-}$ Oxidation number of Cr is 6.

10. **Statement-I** : S_8 disproportionate into S^{2-} and $S_2O_3^{2-}$ in alkaline medium

Statement-II : ClO_4^- undergoes disproportionate in acidic medium.

- (1) Statement I is correct and statement II is incorrect
 (2) Statement I is incorrect and statement II is correct
 (3) Both statement I and statement II are correct
 (4) Both statement I and statement II are incorrect

Ans. (1)

- Sol. (i) $S_8 + OH^- \longrightarrow S^{2-} + S_2O_3^{2-}$
 So S_8 disproportionate in alkaline medium.
 (ii) ClO_4^- do not show disproportionate reaction in any medium.

11. Nessler's reagent is used for identification of following cation:

- (1) Na^+ (2) K^+ (3) NH_4^+ (4) Pb^{+2}

Ans. (3)

- Sol. $NH_4^+ + 2[HgI_4]^{2-} + 4OH^- \rightarrow HgO, Hg(NH_2)I \downarrow + I^- + H_2O$
 Nessler's reagent brown ppt (iodide of millon's base)

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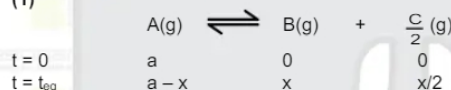
12. For an equilibrium reaction $A(g) \rightleftharpoons B(g) + \frac{C}{2}(g)$ the relation between equilibrium constant (K_p), degree of dissociation (α) and total equilibrium pressure (P) is :

$$(1) K_p = \frac{\alpha^{3/2} \cdot P^{1/2}}{(2+\alpha)^{1/2}(1-\alpha)} \quad (2) K_p = \frac{\alpha^{1/2} \cdot P^{1/2}}{\left(1+\frac{\alpha}{2}\right)^{1/2}(1-\alpha)}$$

$$(3) K_p = \frac{\alpha^{3/2} \cdot P^{1/2}}{\left(1+\frac{\alpha}{2}\right)(1-\alpha)} \quad (4) K_p = \frac{\alpha^{1/2} \cdot P}{\left(1+\frac{\alpha}{2}\right)(1-\alpha)}$$

Ans. (1)

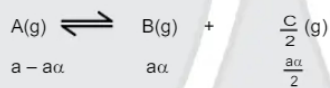
Sol.



For a mole, x moles are dissociated

For 1 mole, $\frac{x}{a}$ moles = α are dissociated

$$x = a\alpha$$



$$\text{Total no. of moles at equilibrium} = a + \frac{a\alpha}{2} = a\left(1 + \frac{\alpha}{2}\right)$$

$$P_{A(g)} = \frac{a(1-\alpha)P}{a\left(1 + \frac{\alpha}{2}\right)} = \frac{(1-\alpha)P}{1 + \frac{\alpha}{2}}$$

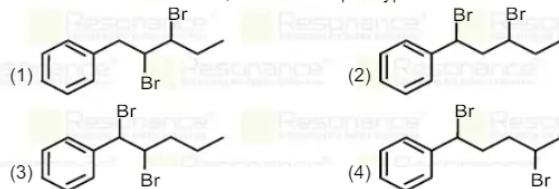
$$P_{B(g)} = \frac{a\alpha \cdot P}{a\left(1 + \frac{\alpha}{2}\right)} = \frac{\alpha P}{1 + \frac{\alpha}{2}}$$

$$P_{C(g)} = \frac{(a\alpha/2) \cdot P}{a\left(1 + \frac{\alpha}{2}\right)} = \frac{(\alpha/2) \cdot P}{1 + \frac{\alpha}{2}}$$

$$K_p = \frac{P_B \cdot (P_C)^{1/2}}{P_A} = \frac{\left(\frac{\alpha}{1 + \frac{\alpha}{2}}\right) \cdot \left(\frac{\alpha/2}{1 + \frac{\alpha}{2}}\right)^{1/2}}{\frac{(1-\alpha)P}{1 + \frac{\alpha}{2}}}$$

$$K_p = \frac{\alpha \cdot \alpha^{1/2} \cdot P^{1/2}}{(2+\alpha)^{1/2}(1-\alpha)} = \frac{\alpha^{3/2} \cdot P^{1/2}}{(2+\alpha)^{1/2}(1-\alpha)}$$

13. Find correct structure of 2,3-Dibromo-1-phenylpentane



Ans. (1)

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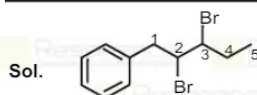
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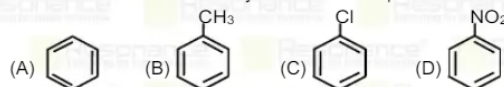
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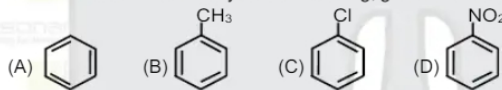
14. The correct order of reactivity towards electrophilic aromatic substitution reaction is.



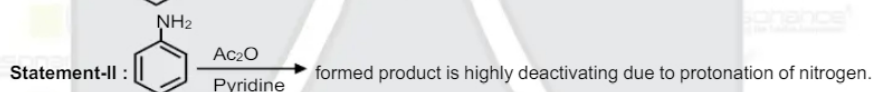
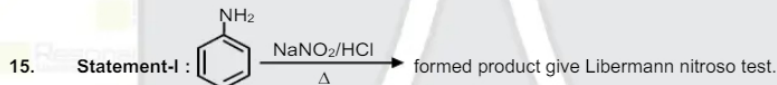
(1) A < B < C < D (2) C < B < D < A (3) D < C < A < B (4) D < C < B < A

Ans. (3)

Sol. Greater the electron density of Aromatic ring, greater will be the rate of electrophilic Aromatic substitution.



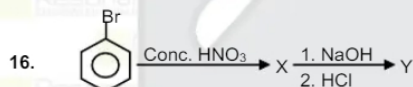
D < C < A < B
Answer is (3).



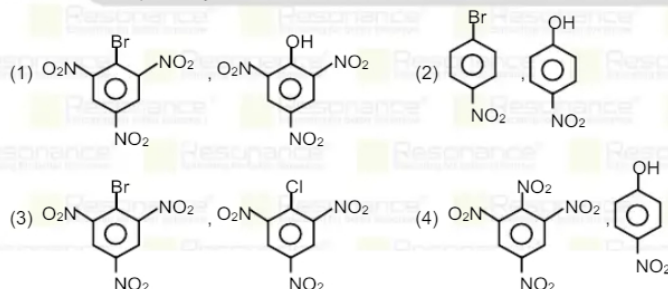
- (1) Both Statement-I & Statement-II are correct.
 (2) Both Statement-I & Statement-II are incorrect.
 (3) Statement-I is correct whereas Statement-II is incorrect.
 (4) Only Statement-II is correct.

Ans. (3)

Sol. In statement-I : Phenol is formed which gives Libermann nitroso test.
 In statement-II : The product benzanilide is weakly activating.



X and Y are respectively.







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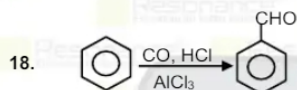
Ans. (1)

17. How many of the vitamins among A, B₁, B₂, B₁₂, C, D and K, can be stored in human body.

- (1) 2 (2) 3 (3) 4 (4) 5

Ans. (2)

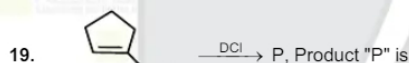
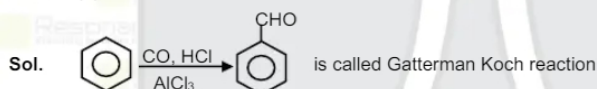
Sol. Only water insoluble and fat soluble vitamins A, D and K can be stored in human body.

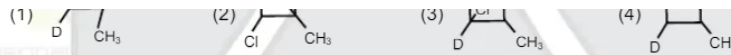


Above reaction is known as

- (1) Etard reaction. (2) Gatterman Koch reaction
 (3) Stephen reaction (4) Rosenmund reaction

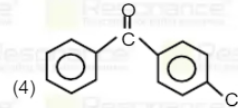
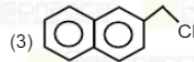
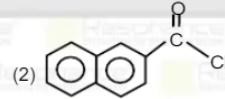
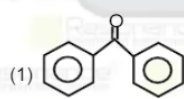
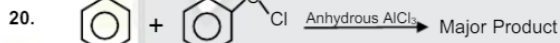
Ans. (2)





Ans. (1)

Sol. It is example of electrophilic addition reaction



Ans. (1)

Sol. It is example of Friedel Craft acylation.

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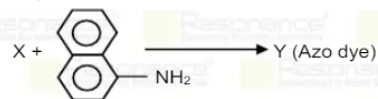
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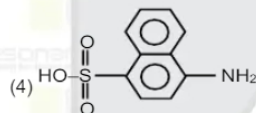
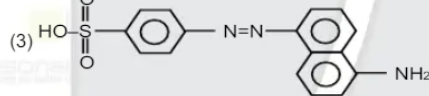
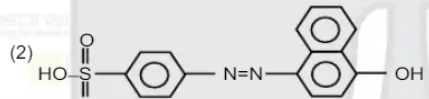
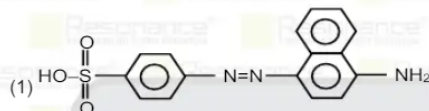
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21. Sulphanic acid + NaNO_2 + CH_3COOH \longrightarrow X



Y is :



Ans. (1)

Sol. Y is red violet Azo dye.

22. Given rate law for a reaction $r = k[A]$

If reaction complete 50% in 120 min then determine in how many minute reaction gets completed 90%?

Ans. (400)

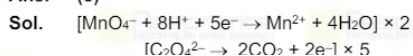
Sol. For first order,

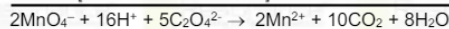
$$t_{90\%} = 3.33 \times t_{50\%}$$

$$= 3.33 \times 120 \approx 400 \text{ min}$$

23. KMnO_4 oxidise $\text{C}_2\text{O}_4^{2-}$ to CO_2 during this reaction no of mole of H^+ ions used with 1 mole of MnO_4^- is

Ans. (8)





24. 1 mole of an ideal gas expands from 10 lit to 100 lit isothermally and reversibly at 300 K, then magnitude of work done is _____ (in kJ) [Nearest integer]

[Given $R = 8.314 \frac{\text{J}}{\text{Mole} \times \text{K}}$]

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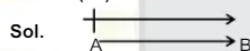
Ans. (6)

Sol. For isothermal reversible process

$$\begin{aligned} W &= -nRT \ln \left(\frac{V_2}{V_1} \right) \\ &= -[1 \times 8.314 \times 300] 2.303 \log \left(\frac{100}{10} \right) \\ &= -2.303 \times 8.314 \times 300 \\ &= -5744 \text{ J} \\ &= -5.744 \text{ kJ} \end{aligned}$$

25. In compound AB dipole moment of A-B bond and bond distance are 1 Å and 1.2 D respectively, then magnitude of fraction of charge on A atom is _____ $\times 10^{-2}$ [Nearest integer]

Ans. (25)

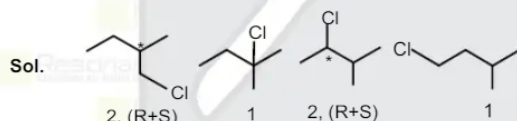


$$\begin{aligned} \mu &= \delta \times d = 1.2 \\ \delta \times 1 \times 10^{-10} &= 1.2 \times 3.33 \times 10^{-30} \text{ C} \times \text{meter} \\ \delta &= 1.2 \times 3.33 \times 10^{-20} \\ &= 3.996 \times 10^{-20} \\ &= 0.3996 \times 10^{-19} \text{ C} \end{aligned}$$

$$\begin{aligned} \text{Fraction of charge} &= \left(\frac{\delta}{e} \right) = \left(\frac{0.3996 \times 10^{-19}}{1.6 \times 10^{-19}} \right) \\ &= 0.24975 = 24.975 \times 10^{-2} \end{aligned}$$

26. Number of monochlorination product of 2-Methylbutane formed in presence of sunlight is.

Ans. (6)

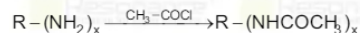


Answer (6)

27. Compound X with molar mass 108 g mol^{-1} undergoes acetylation to give product with molar mass 192 g mol^{-1} . Total number of NH_2 group in benzoid molecule X is.

Ans. (2)

Sol. Let the benzoid molecule has x no of NH_2 group, therefore on acetylation it will form $\text{R}-(\text{NH}-\text{COCH}_3)_x$,



$$\text{R} + 16x = 108$$

$$\text{R} + 58x = 192$$

$$42x = 84$$

$$x = 2.$$

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