

AMRITA VISHVA VIDYAPEETHAM

(University established u/s 3 of UGC Act 1956)

Amrita Entrance Examination – Engineering

PHYSICS, CHEMISTRY & MATHEMATICS

Question Booklet Version Code	D	Question Booklet No.	401750	Time: 2½ Hrs
Number of Pages	20	Number of Questions	100	Max. Marks: 300
Name of the Candidate				
Registration Number				
Signature of the Candidate				

INSTRUCTIONS TO THE CANDIDATES

GENERAL

1. Any malpractice or attempt to commit malpractice in the examination hall will lead to disqualification of the candidate.
2. Candidates are not allowed to carry any textual material, printed or written bits of paper, Mathematical and Physical Tables, Electronic gadgets like tablet, calculator, cell phone, etc. into the examination hall.
3. Candidates shall possess the Amrita Entrance Examination – Engineering Hall Ticket which should be produced on demand.
4. Candidates shall occupy the respective seats bearing their registration numbers.
5. Candidates shall sign the attendance sheet available with the invigilator.
6. **Candidates are not permitted to leave the hall before the end of the examination.**
7. **Candidates are required to handover the ANSWER SHEET and the QUESTION BOOKLET to the invigilator before leaving the hall.**
8. **After submitting the answer sheet, candidates are required to affix their left thumb impression on the attendance sheet available with the invigilator.**

QUESTION BOOKLET

9. **DO NOT OPEN THIS SEALED BOOKLET UNTIL THE INVIGILATOR ANNOUNCES TO DO SO.**
10. **Before opening the Question Booklet**, write the Name, Registration Number and Signature using ball pen in the space provided at the top of this page.
11. **Immediately after opening the booklet**, examine whether it contains all the 100 questions in serial order and 20 pages as mentioned at the top of this page. In case of unprinted, torn or missing pages, the matter should be reported to the invigilator immediately.
12. Rough work may be done on the space provided in this booklet.

(Continued on the last page)

SEAL

Space for rough work

PHYSICS

- An object of size 10 cm is kept at a distance of 10 cm from a convex lens. If the focal length of the lens is 5 cm, the size of the image is
 - 10 cm
 - 20 cm
 - 5 cm
 - 15 cm
- A biconvex lens of focal length 10 cm is to be made from a glass material. If the refractive index of the material is 1.5, what must be the radius of curvature of the surface of the lens?
 - 0.1 m
 - 0.15 m
 - 0.20 m
 - 0.30 m
- A diffraction grating with 10^6 lines / m is used to determine the wavelength of a monochromatic source. The angle of first order diffraction is 30° . The wavelength of the source is
 - 1000 nm
 - 500 nm
 - 400 nm
 - 600 nm
- A glass plate of thickness $1.5 \mu\text{m}$ and refractive index 1.5 is introduced between one of the slits and screen in a Young's double slit experiment. If the wavelength of the monochromatic source used is $\lambda = 0.75 \mu\text{m}$, the phase difference between the interfering waves at the centre of the screen is equal to
 - 6π
 - 3π
 - π
 - 2π
- What is the velocity of light in a medium with refractive index 1.5?
 - 2×10^8 m/s
 - 3×10^8 m/s
 - 1.5×10^8 m/s
 - 2.5×10^8 m/s
- Which among the following electromagnetic radiations is the most energetic?
 - Infra red light
 - Visible light
 - Ultraviolet light
 - microwaves
- Which of the following particles has the shortest de-Broglie wavelength, if all of them move with same speed?
 - beta particle
 - alpha particle
 - proton
 - neutron

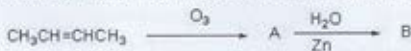
Space for rough work

16. Two forces $F_1 = (7i + 2j)$ N and $F_2 = (-5i + 3j)$ N act on a particle. The third force F_3 that should act on the particle to make it move with constant velocity is
- a) $(2i + 5j)$ N
b) $(-2i - 5j)$ N
c) $(-2i + 5j)$ N
d) $(2i - 5j)$ N
17. Two satellites of masses $3M$ and M orbit the earth in circular orbits of radii r and $3r$ respectively. The ratio of their speeds is
- a) 1 : 1
b) $\sqrt{3}$: 1
c) 3 : 1
d) 9 : 1
18. In an adiabatic process, the pressure of a gas is proportional to the cube of its absolute temperature. The value of γ (which equals C_p/C_v) is
- a) 5/4
b) 4/3
c) 5/3
d) 3/2
19. A mass is moving towards the origin along the x-axis with constant velocity. Its angular momentum with respect to the origin
- a) remains constant
b) is zero
c) increases
d) decreases
20. The rate of cooling of a liquid is 4°C/s , when its temperature is 80°C and is 2°C/s when its temperature is 50°C . The temperature of the surroundings is
- a) 30°C
b) 20°C
c) 10°C
d) 25°C
21. A Charged sphere of radius 1m carries a charge of $1 \times 10^{-9}\text{C}$. The electric fields at a point P, which is at a distance $d = 3\text{m}$ from the centre of the sphere and at a point Q, at a distance $d = 0.3\text{m}$ from the centre of the sphere are respectively
- a) 1 N/C and 100 N/C
b) 1 N/C and zero
c) zero and 1 N/C
d) 1 N/C and 3 N/C
22. An electric dipole lying along X-axis with moment 5Am^2 is subjected to an electric field of magnitude 10J N/C . The torque experienced is
- a) 2 Nm
b) 10 Nm
c) 50 Nm
d) 25 Nm

Space for rough work

CHEMISTRY

31. S_N1 reaction is favored by
- non polar solvents
 - more number of alkyl group on the carbon atom attached to the halogen atom
 - small groups on the carbon attached to the halogen atom
 - no groups on the carbon attached to the halogen atom
32. Phenol is less acidic than
- ethanol
 - o-nitrophenol
 - o-methylphenol
 - o-methoxyphenol
33. Chloro ethane reacts with compound Z to form diethyl ether. Identify Z?
- NaOH
 - H_2SO_4
 - C_2H_5ONa
 - $Na_2S_2O_3$
34. Which of the following reagents may be used to distinguish between phenol and benzoic acid?
- Tollens' reagent
 - Molisch reagent
 - Neutral $FeCl_3$
 - Aqueous NaOH
35. In the following sequence of reactions, the alkene affords the compound 'B'.



The compound B is

- CH_3CHO
 - CH_3COCH_3
 - CH_3CH_2CHO
 - $CH_3CH_2COCH_3$
36. How many chiral carbons are there in β -D-(+)-glucose?
- five
 - six
 - three
 - four
37. Why are certain rubbers called as 'vulcanized rubber'?
- They are formed under volcanic eruption
 - They are prepared by adding 5% of sulphur as cross-linking agent
 - They do not use any co-monomer
 - By the addition of excessive co-monomer

Space for rough work

38. One of the common components of photochemical smog is
- a) formaldehyde
 - b) acetaldehyde
 - c) methane
 - d) CO_2
39. Sodium dodecylbenzenesulphonate refers to
- a) anionic detergent
 - b) soap
 - c) cationic detergent
 - d) nonionic detergent
40. Which one of the following acts as antihistamine?
- a) Equanil
 - b) Morphine
 - c) Serotonine
 - d) Bromophenylamine
41. The actual atomic weight of an element is represented in
- a) number
 - b) "u"
 - c) "amu"
 - d) "mu"
42. The weight of nascent oxygen in milligrams obtained from 6.32 g of potassium permanganate (Molecular weight 158) in acid medium is
- a) 16
 - b) 0.016
 - c) 0.16
 - d) 1.6
43. The value of Plank's constant in units of Js is
- a) 6.626×10^{-34}
 - b) 6.626×10^{-23}
 - c) 6.626×10^{-27}
 - d) 1.38×10^{-23}
44. The mass of proton having a wavelength of 4.2\AA is
- a) 4.78×10^{-33} kg
 - b) 4.78×10^{-33} g
 - c) 7.17×10^{-33} kg
 - d) 2.39×10^{-33} g
45. The measurement of a thermodynamic property known as temperature is based on
- a) zeroth law of thermodynamics
 - b) first law of thermodynamics
 - c) second law of thermodynamics
 - d) kirchoffs equation
46. The bond dissociation enthalpies of $\text{H}_2(\text{g})$, $\text{Cl}_2(\text{g})$ and $\text{HCl}(\text{g})$ are 435, 243 and 431 kJ/mol respectively. The enthalpy of formation of $\text{HCl}(\text{g})$ in kJ/mol will be
- a) 121
 - b) -1211
 - c) -121
 - d) -242
47. Defective coating of zinc over mild steel leads to
- a) enhanced corrosion of mild steel
 - b) increase of corrosion potential
 - c) corrosion of zinc coating
 - d) hydrogen evolution over mild steel

48. What will happen to the rate constant of a reaction when the temperature is raised by 10°C ?
- a) Increase by 10 times
b) Is halved
c) Is doubled
d) Not affected
49. The equivalent conductances at infinite dilution (λ^∞) of ammonium chloride, sodium hydroxide and sodium chloride are 120, 240 and $150\text{ mho cm}^2\text{eq}^{-1}$. The λ^∞ of ammonium hydroxide in $\text{mho cm}^2\text{eq}^{-1}$ is
- a) 270
b) 210
c) 30
d) 510
50. 100 cm^3 of an aqueous solution of protein contains 0.63 g of protein. If the osmotic pressure of the solution at 300K is 2.57×10^{-3} bar, the molar mass of the protein will be
- a) 60039
b) 61039
c) 62039
d) 63039
51. A compound formed by elements P and Q crystallizes in cubic structure in which atoms of P are at corners and atoms of Q are at the face center. The formula of the compound is
- a) AB_3
b) AB
c) A_3B
d) A_3B
52. Syn gas is a mixture of
- a) carbon dioxide and hydrogen
b) carbon monoxide and hydrogen
c) methane and hydrogen
d) methane and carbon monoxide
53. Which one of the following alkali metal hydrides is thermally stable?
- a) Lithium hydride
b) Sodium hydride
c) Potassium hydride
d) Rubidium hydride
54. The correct order of acidic character of the following is
- a) $\text{SO}_2 > \text{CO}_2 > \text{CO} > \text{N}_2\text{O}_5$
b) $\text{SO}_2 > \text{N}_2\text{O}_5 > \text{CO} > \text{CO}_2$
c) $\text{N}_2\text{O}_5 > \text{SO}_2 > \text{CO} > \text{CO}_2$
d) $\text{N}_2\text{O}_5 > \text{SO}_2 > \text{CO}_2 > \text{CO}$
55. Bell metal is an alloy of
- a) copper and tin
b) silver and copper
c) copper and nickel
d) copper, zinc and tin
56. Ammonium dichromate is used in fireworks. The green coloured powder blown in the air is
- a) CrO_3
b) Cr_2O_3
c) Cr
d) $\text{CrO}(\text{O}_2)$

Space for rough work

57. Which one of the following complexing agents is used for the estimation of hardness of water?

- a) Cyanide
- b) Pyrophosphate
- c) EDTA
- d) Ethylene diamine

58. How many σ and π bonds are present in nitromethane

- a) 6 σ and 1 π
- b) 5 σ and 2 π
- c) 6 σ and 2 π
- d) 5 σ and 1 π

59. Retardation factor is calculated as

- a) ratio between 'distance travelled by the substance from the base line and distance moved by the solvent from the base line'
- b) ratio between 'distance travelled by the solvent from the base line and distance moved by the substance from the base line'
- c) sum of 'distance travelled by the substance from the base line and distance moved by the solvent from the base line'
- d) difference of 'distance travelled by the substance from the base line and distance moved by the solvent from the base line'

60. In which one of the following, Mn exhibits its highest oxidation state?

- a) MnO_2
- b) MnO_4^{2-}
- c) MnO_4^-
- d) MnO

Space for rough work

66. Which one of the following is TRUE for any x

a) $\frac{1}{x+5} < \frac{1}{x+2} < \frac{1}{x+3}$

b) $\frac{1}{x+2} < \frac{1}{x+3} < \frac{1}{x+5}$

c) $\frac{1}{x+5} < \frac{1}{x+3} < \frac{1}{x+2}$

d) $\frac{1}{x+3} < \frac{1}{x+2} < \frac{1}{x+5}$

67. The order and degree of the differential equation $y - x \frac{dy}{dx} = \frac{a \frac{dy}{dx}}{\sqrt{1 + (\frac{dy}{dx})^2}}$ is

a) 1, 2

b) 1, 4

c) 1, $5\sqrt{2}$

d) 1, 3

68. The general solution of the differential equation $(1 + e^{(x/y)}) dx + e^{(x/y)} (1 - (x/y)) dy = 0$ is

a) $y + xe^{(x/y)} = C$

b) $x + ye^{(x/y)} = C$

c) $x + C = ye^{(x/y)}$

d) $y + ye^{(x/y)} = C$

69. The triangle with vertices A = (2, 7), B = (4, y) and C = (-2, 6) is right angled at B if the value of y is

a) 10 or -3

b) -10 or -3

c) 10 or 3

d) 9 or 4

70. The point equidistant from the three lines $x + y = 1$, $y = 1$ and $x = 1$ is

a) $(-\frac{1}{\sqrt{2}}, -\frac{1}{\sqrt{2}})$

b) $(+\frac{1}{\sqrt{2}}, -\frac{1}{\sqrt{2}})$

c) $(+\frac{1}{\sqrt{3}}, -\frac{1}{\sqrt{2}})$

d) $(+\frac{1}{\sqrt{2}}, -\frac{1}{\sqrt{3}})$

Space for rough work

87. The equation of the plane through the intersection of the planes $2x - y + z = 6$ and $x + y + 2z = 7$ and passing through the point $(1, 1, 1)$ is

a) $2x - 7y - 5z + 10 = 0$

b) $2x - 7y + 5z + 10 = 0$

c) $2x - 7y - 5z - 10 = 0$

d) $2x + 7y - 5z - 10 = 0$

88. The equation of the line passing through the point $(1, 1, 0)$ and parallel to the plane $3x + 2y + z = 5$ is

a) $\frac{x-1}{-3} = \frac{y-1}{-2} = \frac{z}{1}$

b) $\frac{x+1}{3} = \frac{y+1}{2} = \frac{z}{1}$

c) $\frac{x-1}{3} = \frac{y-1}{2} = \frac{z}{1}$

d) $\frac{x-3}{1} = \frac{y-2}{1} = \frac{z-1}{0}$

89. The angle between the complex numbers $2 + 2i$ and $-7i$ is

a) $\pi/2$

b) $\pi/4$

c) $3\pi/2$

d) $3\pi/4$

90. What is the value of $4 + 5\left(-\frac{1}{2} + i\frac{\sqrt{3}}{2}\right)^{334} + 3\left(-\frac{1}{2} + i\frac{\sqrt{3}}{2}\right)^{365}$?

a) i

b) $\frac{\sqrt{3}}{2}$

c) $\frac{\sqrt{3}}{2}i$

d) $\sqrt{3}i$

Space for rough work

96. The equation $4x^2 + 7y^2 + 32x - 56y + 148 = 0$ represents

- a) an ellipse with center (4, -4)
- b) an ellipse with center (-4, 4)
- c) an ellipse with center (2, -2)
- d) an ellipse with center (-2, 2)

97. The equation for the circle obtained by shifting the circle $x^2 + y^2 = 49$ to 3 units down and 2 units left is:

- a) $(x+3)^2 + (y+2)^2 = 49$
- b) $(x-3)^2 + (y-2)^2 = 49$
- c) $(x-2)^2 + (y-3)^2 = 49$
- d) $(x+2)^2 + (y+3)^2 = 49$

98. The variance of a data set is k , then the variance of the data set obtained by shifting the original data to 3 units is

- a) $k - 3$
- b) $k + 3$
- c) k
- d) $3k$

99. Suppose that $P(A/B) = 0.7$, $P(A) = 0.5$ and $P(B) = 0.2$ then $P(B/A)$ is,

- a) 0.14
- b) 0.4
- c) 0.3
- d) 0.28

100. A medical test is capable of identifying someone with the illness as positive is 99% and someone without illness as negative 95%. If the illness is present in the general population with probability 0.0001, the probability for anyone to have illness when the medical test results positive is

- a) 0.00009
- b) 0.002
- c) 0.0001
- d) 0.9980

Space for rough work

Space for rough work

QUESTION BANK

UNIT - 1: Introduction to Chemistry

1. Matter is anything that has mass and occupies space. It is made up of particles which are continuously moving. Matter can be classified into three states: solid, liquid, and gas. Solids have a definite shape and volume, liquids have a definite volume but no definite shape, and gases have neither a definite shape nor a definite volume.

2. The particles of matter are held together by forces of attraction. These forces are stronger in solids and weaker in gases. The particles in solids are packed closely together, while the particles in gases are far apart and move freely.

3. The change of state of matter is called a physical change. It does not involve the formation of a new substance. Examples include melting, boiling, condensation, and sublimation.

4. The change of state of matter is accompanied by the absorption or release of heat. For example, melting and boiling require heat, while condensation and freezing release heat.

5. The boiling point of a liquid is the temperature at which it changes from a liquid to a gas. The boiling point of water is 100°C at standard atmospheric pressure.

6. The melting point of a solid is the temperature at which it changes from a solid to a liquid. The melting point of ice is 0°C at standard atmospheric pressure.

7. The sublimation of a solid is the change of state from a solid to a gas without passing through the liquid state. Examples include the sublimation of iodine and naphthalene.

8. The particles of matter are in constant motion. The kinetic energy of the particles increases with temperature. In solids, the particles vibrate about their fixed positions. In liquids, they move more freely, and in gases, they move rapidly in all directions.

9. The rate of evaporation of a liquid depends on several factors, including temperature, surface area, and wind speed. Higher temperature, larger surface area, and higher wind speed all increase the rate of evaporation.

10. The boiling point of a liquid decreases with increasing altitude. This is because the atmospheric pressure is lower at higher altitudes, and the boiling point of a liquid is lower at lower pressures.

(continued from the first page)

OMR ANSWER SHEET

13. Use the OMR answer sheet carefully; no spare sheet will be issued under any circumstance.
14. Do not fold or make any stray mark on the OMR sheet.
15. Use HB Pencil or Black ball point pen for shading the bubbles and ball point pen for writing.
16. In the OMR answer sheet, make the following entries
 - a. Write the Registration Number, Question Booklet Number and Question Booklet Version code using ball point pen.
 - b. Fill the ovals corresponding to the Registration Number, Question Booklet Number and Question Booklet Version Code using HB pencil / ball point pen.
 - c. Write your Name and Signature using ball point pen.
17. Rough work should not be done on the answer sheet.

ANSWERING AND EVALUATION

18. For each question, four answers are suggested of which only one is correct / most appropriate. Mark the correct / most appropriate answer by darkening the corresponding bubble using HB pencil or Black ball point pen.
19. In case the candidate wishes to change the choice already shaded using HB pencil, he/she may erase the marking completely and thereafter shade the alternative bubble. If ball point pen is used for shading the ovals, make sure of the answer before shading since such markings cannot be altered.
20. If more than one bubble is darkened against a question, it will be treated as an incorrect answer.
21. For each correct answer, three marks will be awarded.
22. **For each incorrect answer, one mark will be deducted from the total score.**
23. If any smudge is left on the OMR sheet, evaluation will become imperfect.

SEAL