

Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

The syllabus is in tune with the syllabus adapted by the Board of Intermediate Education TS, for Intermediate course with effect from the academic year 2020-2021 (1st year) (70%) and 2021-2022 (2nd year) (70%) and is designed at the level of Intermediate Course and equivalent to (10+2) scheme of Examination conducted by Board of Intermediate Education, TS. The syllabus is designed to indicate the scope of subjects included for TS EAMCET-2022. The topics mentioned therein are not to be regarded as exhaustive. Questions may be asked in TS EAMCET-2022 to test the student's knowledge and intelligent understanding of the subject. The syllabus is applicable to students of both the current and previous batches of Intermediate Course, who desire to appear for TS EAMCET-2022.

MATHEMATICS

1) ALGEBRA

- a) **Functions:** Types of functions – Definitions – Domain, Range.
- b) **Matrices:** Types of matrices - Scalar multiple of a matrix and multiplication of matrices - Transpose of a matrix – Determinants (up to 3.4.7) - Adjoint and Inverse of a matrix - Solution of simultaneous linear equations – Cramer's Rule, matrix inversion method.
- c) **Complex Numbers:** Complex number as an ordered pair of real numbers - fundamental operations - Representation of complex numbers in the form $a + ib$.
- d) **De Moivre's Theorem:** De Moivre's theorem- Integral and Rational indices - n^{th} roots of unity- Geometrical Interpretations – Illustrations. (Exercise 2(b) Section II and III are deleted)
- e) **Quadratic Expressions:** Quadratic expressions, equations in one variable - Sign of quadratic expressions – Change in signs – Maximum and minimum values.
- f) **Theory of Equations:** The relation between the roots and coefficients in an equation - Solving the equations when two or more roots of it are connected by certain relation - Equation with real coefficients, occurrence of complex roots in conjugate pairs and its consequences - Transformation of equations - Reciprocal Equations.
- g) **Permutations and Combinations:** Fundamental Principle of counting – linear permutations- Permutations of 'n' dissimilar things taken 'r' at a time – Combinations - definitions, certain theorems.
- h) **Binomial Theorem:** Binomial theorem for positive integral index - Binomial theorem for rational Index (without proof).
- i) **Partial fractions:** Rational functions - Partial fractions - Partial fractions of $f(x)/g(x)$ when $g(x)$ contains non-repeated linear factors - Partial fractions of $f(x)/g(x)$ when $g(x)$ contains repeated and/or non-repeated linear factors - Partial fractions of $f(x)/g(x)$ when $g(x)$ contains irreducible factors only.

2) TRIGONOMETRY:

- a) **Trigonometric Ratios up to Transformations:** Graphs and Periodicity of Trigonometric functions - Trigonometric ratios and Compound angles - Trigonometric

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ratios of multiple and sub- multiple angles - Transformations - Sum and Product rules.

- b) **Hyperbolic Functions:** Definition of Hyperbolic Function – Graphs - Addition formulae of Hyperbolic Functions.
- c) **Properties of Triangles:** Relation between sides and angles of a Triangle - Sine, Cosine, Tangent and Projection rules- Half angle formulae and areas of a triangle – Incircle and Excircle of a Triangle.
- 3) VECTOR ALGEBRA:**
- a) **Addition of Vectors:** Vectors as a triad of real numbers - Classification of vectors - Addition of vectors - Scalar multiplication - Angle between two non-zero vectors- Linear combination of vectors - Component of a vector in three dimensions - Vector equations of line and plane including their Cartesian equivalent forms.
- b) **Product of Vectors:** Scalar Product - Geometrical Interpretations - orthogonal projections - Properties of dot product - Expression of dot product in i, j, k system - Angle between two vectors - Geometrical Vector methods - Vector equations of plane in normal form - Angle between two planes - Vector product of two vectors and properties - Vector product in i, j, k system - Vector Areas - Scalar Triple Product.
- 4) PROBABILITY:**
- a) **Measures of Dispersion:** Range - Mean deviation for ungrouped data.
- b) **Probability:** Random experiments and events - Classical definition of probability, Axiomatic approach and addition theorem of probability - Independent and dependent events - conditional probability - multiplication theorem.
- c) **Random Variables and Probability Distributions:** Random Variables - Theoretical discrete distributions – Binomial and Poisson Distributions.
- 5) COORDINATE GEOMETRY:**
- a) **Locus:** Definition of locus – Illustrations - To find equations of locus - Problems connected to it.
- b) **Transformation of Axes:** Transformation of axes - Rules, Derivations and Illustrations - Rotation of axes - Derivations – Illustrations.
- c) **The Straight Line:** Revision of fundamental results - Straight line - Normal form – Illustrations - Straight line - Symmetric form - Straight line - Reduction into various forms - Intersection of two Straight Lines - Family of straight lines - Concurrent lines - Condition for Concurrent lines - Angle between two lines - Length of perpendicular from a point to a Line - Distance between two parallel lines- Concurrent lines - properties related to a triangle.
- d) **Pair of Straight lines:** Equations of pair of lines passing through origin - angle between a pair of lines - Condition for perpendicular and coincident lines - Pair of lines - second degree general equation - Homogenizing a second degree equation with a first degree equation in x and y .



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- e) **Circle:** Equation of circle - standard form - center and radius of a circle with a given line segment as diameter & equation of circle through three non collinear points - parametric equations of a circle - Position of a point in the plane of a circle – power of a point - definition of tangent - length of tangent - Position of a straight line in the plane of a circle - conditions for a line to be tangent – chord joining two points on a circle – equation of the tangent at a point on the circle - point of contact-equation of normal - Chord of contact - pole and polar - conjugate points and conjugate lines - equation of chord with given middle point - Relative position of two circles - circles touching each other externally, internally - common tangents – centers of similitude .
- f) **System of circles:** Angle between two intersecting circles - Radical axis of two circles - properties - Common chord and common tangent of two circles – radical center - Intersection of a line and a Circle.
- g) **Parabola:** Conic sections – Parabola - equation of parabola in standard form - different forms of parabola - parametric equations.
- h) **Ellipse:** Equation of ellipse in standard form - Parametric equations.
- i) **Hyperbola:** Equation of hyperbola in standard form- Parametric equations.
- j) **Three Dimensional Coordinates:** Coordinates - Section formulae - Centroid of a triangle and tetrahedron.
- k) **Direction Cosines and Direction Ratios:** Direction Cosines - Direction Ratios.
- l) **Plane:** Cartesian equation of Plane - Simple Illustrations.
- 6) **CALCULUS:**
 - a) **Limits and Continuity:** Intervals and neighbourhoods – Limits - Standard Limits.
 - b) **Differentiation:** Derivative of a function - Elementary Properties - Trigonometric, Hyperbolic - Derivatives.
 - c) **Applications of Derivatives:** Errors and approximations - Geometrical Interpretation of a derivative - Equations of tangents and normal - Angles between two curves and condition for orthogonality of curves - Maxima and Minima.
 - d) **Integration:** Integration as the inverse process of differentiation - Standard forms - properties of integrals - Method of substitution - integration of Algebraic, exponential, logarithmic and trigonometric.
 - e) **Definite Integrals:** Definite Integral as the limit of sum - Interpretation of Definite Integral as an area - Fundamental theorem of Integral Calculus – Properties.
 - f) **Differential equations:** Formation of differential equation - Degree and order of an ordinary differential equation - Solving differential equation by Variables separable method.

PHYSICS

- 1) PHYSICAL WORLD:** What is physics? Fundamental forces in nature.
- 2) UNITS AND MEASUREMENTS:** Introduction, The international system of units, Measurement of Length, Measurement of Large Distances, Estimation of Very Small Distances: Size of a Molecule, Range of Lengths, Measurement of Mass, Range of Masses, Measurement of time, Accuracy, precision of instruments and errors in measurement, Systematic errors, random errors, least count error, Absolute Error, Relative Error and Percentage Error, Combination of Errors, Significant figures, Rules for Arithmetic Operations with Significant Figures, Rounding off the Uncertain Digits, Rules for Determining the Uncertainty in the Results of Arithmetic Calculations, Dimensions of Physical Quantities, Dimensional Formulae and dimensional equations, Dimensional Analysis and its Applications, Checking the Dimensional Consistency of Equations, Deducing Relation among the Physical Quantities.
- 3) MOTION IN A STRAIGHT LINE:** Introduction, Position, path length and displacement, Average velocity and average speed, Instantaneous velocity and speed, Acceleration. Kinematic equations for uniformly accelerated motion, Relative velocity.
- 4) MOTION IN A PLANE:** Introduction, Scalars and vectors, Position and Displacement Vectors, Equality of Vectors, Multiplication of vectors by real numbers, Addition and subtraction of vectors - graphical method, Resolution of vectors, Vector addition - analytical method, Motion in a plane, Position Vector and Displacement, Velocity, Acceleration, Motion in a plane with constant acceleration, Relative velocity in two dimensions. Projectile motion, Equation of path of a projectile, Time of Maximum height, Maximum height of a projectile, Horizontal range of projectile, Uniform circular motion.
- 5) LAWS OF MOTION:** Introduction, Conservation of momentum, Equilibrium of a particle, Common forces in mechanics, friction, Circular motion, Motion of a car on a level road, Motion of a car on a Banked Road, Solving problems in mechanics.
- 6) WORK, ENERGY AND POWER:** Introduction, The Scalar Product, Notions of work and kinetic energy : The work-energy theorem, Work, Kinetic energy, Work done by a variable force, The work-energy theorem for a variable force, The concept of Potential Energy, The conservation of Mechanical Energy, The Potential Energy of a spring, Various forms of energy: the law of conservation of energy, Heat, Chemical Energy, Electrical Energy, The Equivalence of Mass and Energy, Nuclear Energy.

The Principle of Conservation of Energy, Power, Collisions, Elastic and Inelastic Collisions, Collisions in one dimension, Coefficient of Restitution and its determination, Collisions in Two Dimensions.
- 7) SYSTEMS OF PARTICLES AND ROTATIONAL MOTION:** Introduction, What kind of motion can a rigid body have?, Centre of mass, Centre of Gravity, Motion of centre of mass, Linear momentum of a system of particles, Vector product of two vectors, Angular velocity and its relation with linear velocity, Angular acceleration, Kinematics of rotational motion about a fixed axis, Torque and angular momentum, Moment of force (Torque), Angular momentum of particle, Torque and angular momentum for a system of a particles, conservation of angular momentum, Equilibrium of a rigid body, Principle of moments,



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Moment of inertia, Dynamics of rotational motion about a fixed axis, Angular momentum in case of rotation about a fixed axis, Conservation of Angular Momentum, Rolling motion, Kinetic Energy of Rolling Motion.

8) OSCILLATIONS: Introduction, Periodic and oscillatory motions, Period and frequency, Displacement, Simple harmonic motion (S.H.M.), Simple harmonic motion and uniform circular motion, Velocity and acceleration in simple harmonic motion, Force law for Simple harmonic Motion, Energy in simple harmonic motion, Some systems executing Simple Harmonic Motion, Oscillations due to a spring, The Simple Pendulum, Damped simple harmonic motion, Forced oscillations and resonance.

9) GRAVITATION: Introduction, Universal law of gravitation, Acceleration due to gravity below and above the surface of earth, Gravitational potential energy, Escape speed, Earth satellite, Energy of an orbiting satellite, Geostationary and polar satellites, Weightlessness.

10) MECHANICAL PROPERTIES OF SOLIDS: Introduction, Elastic behaviour of solids, Stress and strain, Hooke's law, Stress-strain curve, Elastic moduli, Young's Modulus, Bulk Modulus.

11) MECHANICAL PROPERTIES OF FLUIDS: Introduction, Pressure, Pascal's Law, Variation of Pressure with Depth, Atmosphere Pressure and Gauge Pressure, Hydraulic Machines, Streamline flow, Bernoulli's principle, Speed of Efflux: Torricelli's Law, Venturi-meter, Blood Flow and Heart Attack, Dynamic Lift.

Viscosity, Variation of Viscosity of fluids with temperature, Stokes' Law, Reynolds number, critical velocity, Surface tension, Surface Energy, Surface Energy and Surface Tension, Angle of Contact, Drops and Bubbles, Capillary Rise, Detergents and Surface Tension.

12) THERMAL PROPERTIES OF MATTER: Introduction, Temperature and heat, Measurement of temperature, Ideal-gas equation and absolute temperature, Thermal expansion. Specific heat capacity, Calorimetry, Change of state, Regelation, Latent Heat, Black body Radiation, Greenhouse Effect, Newton's law of cooling.

13) THERMODYNAMICS: Introduction, Thermal equilibrium, Zeroth law of thermodynamics, Heat, Internal Energy and work, First law of thermodynamics, Specific heat capacity. Thermodynamic state variables and equation of State, Thermodynamic process, Quasi-static process, Isothermal Process, Adiabatic Process, Isochoric Process, Isobaric process, Cyclic process, Second law of thermodynamics, Reversible and irreversible processes.

14) KINETIC THEORY: Introduction, Molecular nature of matter, Behavior of gases, Boyle's Law, Charles' Law, Kinetic theory of an ideal gas, Pressure of an Ideal Gas, kinetic interpretation of temperature, Law of equipartition of energy, Specific heat capacity, Monatomic Gases, Diatomic Gases, Polyatomic Gases, Specific Heat Capacity of Solids, Specific Heat Capacity of Water, Mean free path.



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- 15) **WAVES:** Introduction, Transverse and longitudinal waves, Displacement relation in a progressive wave, The speed of a travelling wave, The principle of superposition of waves, Reflection of waves, Beats.
- 16) **RAY OPTICS AND OPTICAL INSTRUMENTS:** Introduction, Refraction, Total Internal Reflection, Refraction at Spherical Surfaces and by Lenses, Refraction through a Prism, Dispersion by a Prism, Some Natural Phenomena due to Sunlight, The Rainbow, Optical Instruments.
- 17) **WAVE OPTICS:** Introduction, Huygens Principle, Refraction and reflection of plane waves using Huygens Principle, Coherent and Incoherent Addition of Waves, Interference of Light Waves and Young's Experiment, Diffraction, The single slit, seeing the single slit diffraction pattern. The validity of ray optics
- 18) **ELECTRIC CHARGES AND FIELDS:** Introduction, Electric Charges, Conductors and Insulators, Charging by Induction, Basic Properties of Electric Charge, Coulomb's Law, Forces between Multiple Charges, Electric Field, Electric Field Lines, Electric Flux, Electric Dipole, Dipole in a Uniform External Field, Continuous Charge Distribution, Gauss's Law, Application of Gauss's Law, Field due to an infinitely long straight uniformly charged wire, Field due to a uniformly charged infinite plane sheet.
- 19) **ELECTROSTATIC POTENTIAL AND CAPACITANCE:** Introduction, Electrostatic Potential, Potential due to a Point Charge, Potential due to an Electric Dipole, Potential due to a System of Charges, Equipotential Surfaces, Potential Energy of a System of Charges, Potential Energy in an External Field, Electrostatics of Conductors, Dielectrics and Polarisation, Capacitors and Capacitance, The Parallel Plate Capacitor, Effect of Dielectric on Capacitance, Combination of Capacitors, Energy Stored in a Capacitor.
- 20) **CURRENT ELECTRICITY:** Introduction, Electric Current, Electric Currents in Conductors, Ohm's law, Drift of Electrons and the Origin of Resistivity, Limitations of Ohm's Law, Temperature Dependence of Resistivity.
Electrical Energy, Power, Cells, emf, Internal Resistance, Cells in Series and in Parallel, Kirchhoff's rules, Wheatstone Bridge, Meter Bridge, Potentiometer.
- 21) **MOVING CHARGES AND MAGNETISM:** Introduction, Magnetic Force, Motion in a Magnetic Field, Magnetic Field due to a Current Element, Biot-Savart Law, Magnetic Field on the Axis of a Circular Current Loop, Ampere's Circuital Law.
The Solenoid and the Toroid, Force between Two Parallel Currents, the Ampere, Torque on Current Loop, Magnetic Dipole, The Moving Coil Galvanometer.
- 22) **MAGNETISM AND MATTER:** Introduction, The Bar Magnet, The magnetic field lines, The Electrostatic analog, Magnetism and Gauss's Law, The Earth's Magnetism, Magnetisation and Magnetic Intensity.
- 23) **ELECTROMAGNETIC INDUCTION:** Introduction, The Experiments of Faraday and Henry, Magnetic Flux, Faraday's Law of Induction, Lenz's Law and Conservation of Energy, Motional Electromotive Force, Energy Consideration: A Quantitative Study, Eddy Currents, Inductance, AC Generator.
- 24) **ALTERNATING CURRENT:** Introduction, AC Voltage Applied to a Resistor, Representation of AC Current and Voltage by Rotating Vectors — Phasors, AC Voltage



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Applied to an Inductor, AC Voltage Applied to a Capacitor, AC Voltage Applied to a Series LCR Circuit, LC Oscillations, Transformers.

- 25) **ELECTROMAGNETIC WAVES:** Introduction, Electromagnetic Waves, Electromagnetic Spectrum.
- 26) **DUAL NATURE OF RADIATION AND MATTER:** Introduction, Electron Emission, Photoelectric Effect, Experimental Study of Photoelectric Effect, Photoelectric Effect and Wave Theory of Light, Einstein's Photoelectric Equation.
Energy Quantum of Radiation, Particle Nature of Light: The Photon, Wave Nature of Matter.
- 27) **ATOMS:** Introduction, Alpha-particle Scattering and Rutherford's Nuclear Model of Atom, Atomic Spectra, Bohr Model of the Hydrogen Atom, The Line Spectra of the Hydrogen Atom, DE Broglie's Explanation of Bohr's Second Postulate of Quantisation.
- 28) **NUCLEI:** Introduction, Atomic Masses and Composition of Nucleus, Size of the Nucleus, Mass-Energy, Nuclear Force, Nuclear Energy.
- 29) **SEMICONDUCTOR ELECTRONICS: MATERIALS, DEVICES AND SIMPLE CIRCUITS:** Introduction, Classification of Materials: Metals, Semiconductors and Insulators, Intrinsic Semiconductor, Extrinsic Semiconductor, p-n Junction, Semiconductor diode, Application of Junction Diode as a Rectifier, Special Purpose p-n Junction Diodes, optoelectronic junction devices.
Junction Transistor, transistor structure and action, basic transistor circuit configurations and transistor characteristics, Digital Electronics and Logic Gates, Integrated Circuits.
- 30) **COMMUNICATION SYSTEMS:** Introduction, Elements of a Communication System, Basic Terminology Used in Electronic Communication Systems, Bandwidth of Signals, Bandwidth of Transmission Medium, Propagation of Electromagnetic Waves, Modulation and its Necessity, Amplitude Modulation, Production of Amplitude Modulated Wave, Detection of Amplitude Modulated Wave.

CHEMISTRY

1) ATOMIC STRUCTURE:

Developments leading to the Bohr's model of atom; Nature of electromagnetic radiation; Particle nature of electromagnetic radiation- Planck's quantum theory; Bohr's model for Hydrogen atom; Explanation of line spectrum of hydrogen; Limitations of Bohr's model; Quantum mechanical considerations of sub atomic particles; Dual behaviour of matter; Heisenberg's uncertainty principle;

Orbitals and quantum numbers; Energies of orbitals; Filling of orbitals in atoms. Aufbau Principle, Pauli's exclusion Principle and Hund's rule of maximum multiplicity; Electronic configurations of atoms; Stability of half filled and completely filled orbitals.

2) CLASSIFICATION OF ELEMENTS AND PERIODICITY IN PROPERTIES:

Modern periodic law and present form of the periodic table; Nomenclature of elements with atomic number greater than 100; Electronic configuration of elements and the periodic table; Electronic configuration and types of elements s, p, d and f blocks; Trends in physical properties: (a) Atomic radius, (b) Ionic radius (c) Variation of size in inner transition elements (d) Ionization enthalpy, (e) Electron gain enthalpy, (f) Electro negativity; Periodic trends in chemical properties: (a) Periodicity of valence or Oxidation states, (b) Anomalous properties of second period elements - diagonal relationship; Periodic trends and chemical reactivity.

3) CHEMICAL BONDING AND MOLECULAR STRUCTURE:

Kossel - Lewis approach to chemical bonding, Octet rule, covalent bond-Lewis representation of simple molecules, formal charges, limitations of octet rule; Ionic or electrovalent bond - Crystal structure of sodium chloride, Lattice enthalpy; General properties of ionic compounds; Bond Parameters - bond length, bond angle, bond enthalpy, bond order, resonance - Polarity of bonds - dipole moment; Fajan's rules; Valence Shell Electron Pair Repulsion (VSEPR) theory; Predicting the geometry of simple molecules; Valence bond theory-Orbital overlap concept-Directional properties of bonds-overlapping of atomic orbitals, strength of sigma and pi bonds-Factors favouring the formation of covalent bonds; Hybridisation- different types of hybridization involving s, p and d orbitals- shapes of simple covalent molecules; Coordinate bond -definition with examples; Molecular orbital theory - Formation of molecular orbitals, Linear combination of atomic orbitals (LCAO)-conditions for combination of atomic orbitals - Energy level diagrams for molecular orbitals -Bonding in some homo nuclear diatomic molecules- H₂, He₂, Li₂, B₂, C₂, N₂ and O₂; Hydrogen bonding-cause of formation of hydrogen bond - Types of hydrogen bonds-inter and intra molecular-General properties of hydrogen bonds.

4) STATES OF MATTER: GASES AND LIQUIDS:

Intermolecular forces; Thermal Energy; Intermolecular forces vs Thermal interactions; The Gaseous State; The Gas Laws; Ideal gas equation; Graham's law of diffusion; Dalton's Law of partial pressures; Kinetic molecular theory of gases; Kinetic gas equation of an ideal gas (No derivation) deduction of gas laws from Kinetic gas equation; Behaviour of real gases - Deviation from Ideal gas behaviour - Compressibility factor vs Pressure diagrams of real gases.

5) STOICHIOMETRY:

Importance of chemistry-some basic concepts - Properties of matter - uncertainty in measurement-significant figures, dimensional analysis; Atomic and molecular masses- mole concept and molar mass. Concept of equivalent weight; Percentage composition of compounds

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and calculations of empirical and molecular formulae of compounds; Stoichiometry and stoichiometric calculations; Redox reactions-classical idea of redox reactions, oxidation and reduction reactions-redox reactions in terms of electron transfer; Oxidation number concept; Types of Redox reactions-combination, decomposition, displacement and disproportionation reactions; Balancing of redox reactions - oxidation number method, Half reaction (ion-electron) method.

6) THERMODYNAMICS:

Thermodynamic Terms; The system and the surroundings; Types of systems and surroundings; The state of the system; The Internal Energy as a State Function. (a) Work (b) Heat (c) The general case, the first law of Thermodynamics; Applications-Work; Enthalpy- a useful new state function; Extensive and intensive properties; Measurement of ΔU and ΔH : Calorimetry; Enthalpy change, $\Delta_r H$ of reactions - reaction Enthalpy (a) Standard enthalpy of reactions, (b) Enthalpy changes during phase transformations, (c) Standard enthalpy of formation, (d) Thermo chemical equations (e) Hess's law of constant heat summation; Enthalpies for different types of reactions. (a) Standard enthalpy of combustion ($\Delta_c H^\theta$), (b) Enthalpy of atomization ($\Delta_a H^\theta$), phase transition, sublimation and ionization, (c) Bond Enthalpy ($\Delta_{\text{bond}} H^\theta$), (d) Enthalpy of solution ($\Delta_{\text{sol}} H^\theta$) and dilution; Lattice enthalpy Spontaneity. (a) Is decrease in enthalpy a criterion for spontaneity? (b) Entropy and spontaneity, the second law of thermodynamics, (c) Gibbs Energy and spontaneity; Absolute entropy and the third law of thermodynamics.

7) CHEMICAL EQUILIBRIUM AND ACIDS-BASES:

Equilibrium in Physical processes; Equilibrium in chemical processes - Dynamic Equilibrium; Law of chemical Equilibrium - Law of mass action and Equilibrium constant; Homogeneous Equilibria, Equilibrium constant in gaseous systems. Relationship between K_p and K_c ; Heterogeneous Equilibria; Applications of Equilibrium constant; Factors affecting Equilibria - Le Chatelier's principle application to industrial synthesis of Ammonia and Sulphur trioxide; Ionic Equilibrium in solutions; Acids, bases and salts- Arrhenius, Bronsted-Lowry and Lewis concepts of acids and bases; Buffer solutions-designing of buffer solution-Preparation of Acidic buffer; Solubility Equilibria of sparingly soluble salts. Solubility product constant, Common ion effect on solubility of ionic salts.

8) HYDROGEN AND ITS COMPOUNDS:

Position of hydrogen in the periodic table; Dihydrogen-Occurrence and Isotopes; Hydrides: Ionic, covalent, and non-stoichiometric hydrides; Water: Physical properties; structure of water, ice. Chemical properties of water; hard and soft water, temporary and permanent hardness of water; Heavy Water; Hydrogen as a fuel.

9) s - BLOCK ELEMENTS (ALKALI AND ALKALINE EARTH METALS)

Group 1 Elements : Alkali metals; Electronic configurations; Atomic and Ionic radii; Ionization enthalpy; Hydration enthalpy; Physical properties; Chemical properties; Uses; General characteristics of the compounds of the alkali metals: Oxides; Halides; Salts of oxo acids; Anomalous properties of Lithium: Differences and similarities with other alkali metals, Diagonal relationship; similarities between Lithium and Magnesium;

Group 2 Elements: Alkaline earth elements; Electronic configuration; Ionization enthalpy; Hydration enthalpy; Physical properties, Chemical properties; Uses; General characteristics of compounds of the Alkaline Earth Metals: Oxides, hydroxides, halides, salts of oxoacids (Carbonates; Sulphates and Nitrates); Anomalous behavior of Beryllium; its diagonal relationship with Aluminium.



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10) p- BLOCK ELEMENTS GROUP 13 (BORON FAMILY):

General introduction - Electronic configuration, Atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties; Important trends and anomalous properties of boron.

11) p-BLOCK ELEMENTS - GROUP 14 (CARBON FAMILY):

General introduction - Electronic configuration, Atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties; Important trends and anomalous properties of carbon; Allotropes of carbon; Uses of carbon;

12) ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES, TECHNIQUES AND HYDROCARBONS SOME BASIC PRINCIPLES AND TECHNIQUES

General introduction; Tetravalency of Carbon: shapes of organic compounds; Structural representations of organic compounds; Classification of organic compounds; Nomenclature of organic compounds; Isomerism; Fundamental concepts in organic reaction mechanisms; Fission of covalent bond; Nucleophiles and electrophiles; Electron movements in organic reactions; Electron displacement effects in covalent bonds: inductive effect, resonance, resonance effect, electromeric effect, hyperconjugation; Types of Organic reactions;

HYDROCARBONS

Classification of Hydrocarbons; **Alkanes** - Nomenclature, isomerism (structural and conformations of ethane only); Preparation of alkanes; Properties - Physical properties and chemical reactivity, Controlled Oxidation, Isomerisation, Aromatization, reaction with steam; **Alkenes**- Nomenclature, structure of ethene, Isomerism (structural and geometrical); Methods of preparation; Properties- Physical and chemical reactions: Addition of hydrogen, halogen, water, sulphuric acid, hydrogen halides (Mechanism- ionic and peroxide effect, Markovnikov's, anti Markovnikov's or Kharasch effect). Oxidation, Ozonolysis and Polymerization;

Alkynes - Nomenclature and isomerism, structure of acetylene. Methods of preparation of acetylene; Physical properties, Chemical reactions- acidic character of acetylene, addition reactions- of hydrogen, Halogen, Hydrogen halides and water. Polymerization; **Aromatic Hydrocarbons**: Nomenclature and isomerism, Structure of benzene, Resonance and aromaticity; Preparation of benzene. Physical properties. Chemical properties: Mechanism of electrophilic substitution. Electrophilic substitution reactions- Nitration, Sulphonation, Halogenation, Friedel-Crafts alkylation and acylation; Directive influence of functional groups in mono substituted benzene; Carcinogenicity and toxicity

13) SOLID STATE:

General characteristics of solids state; Amorphous and crystalline solids; Classification of crystalline solids based on different binding forces (molecular, ionic, metallic and covalent solids); Probing the structure of solids: X-ray crystallography; Crystal lattices and unit cells. Number of atoms in a unit cell (primitive, body centred and face centred cubic unit cell); Close packed structures: Close packing in one dimension, in two dimensions and in three dimensions- tetrahedral and octahedral voids- formula of a compound and number of voids filled- locating tetrahedral and octahedral voids; Packing efficiency in simple cubic, bcc, hcp, ccp lattice; Calculations involving unit cell dimensions-density of the unit cell; Imperfections in solids-types of point defects-stoichiometric and non-stoichiometric defects.



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14) SOLUTIONS:

Types of solutions; Expressing concentration of solutions - mass percentage, volume percentage, mass by volume percentage, parts per million, mole fraction, molarity and molality; Solubility: Solubility of a solid in a liquid, solubility of a gas in a liquid, Henry's law; Vapour pressure of liquid solutions: vapour pressure of liquid- liquid solutions. Raoult's law as a special case of Henry's law -vapour pressure of solutions of solids in liquids; Ideal and non-ideal solutions; Colligative properties and determination of molar mass-relative lowering of vapour pressure-elevation of boiling point-depression of freezing point-osmosis and osmotic pressure-reverse osmosis and water purification.

15) ELECTROCHEMISTRY AND CHEMICAL KINETICS:

ELECTROCHEMISTRY:

Nernst equation-equilibrium constant from Nernst equation- electro chemical cell and Gibbs energy of the cell reaction; Conductance of electrolytic solutions- measurement of the conductivity of ionic solutions-variation of conductivity and molar conductivity with concentration-strong electrolytes and weak electrolytes-applications of Kohlrausch law; Electrolysis: Faraday's laws of electrolysis-products of electrolysis.

CHEMICAL KINETICS: Rate of a chemical reaction; Factors influencing rate of a reaction: dependance of rate on concentration- rate expression and rate constant- order of a reaction, molecularity of a reaction; Integrated rate equations-zero order reactions-first order reactions- half life of a reaction; Pseudo first order reaction; Temperature dependence of the rate of a reaction. Effect of catalyst.

16) SURFACE CHEMISTRY:

Adsorption and absorption: Distinction between adsorption and absorption-mechanism of adsorption-types of adsorption-characteristics of physisorption-characteristics of chemisorptions-adsorption isotherms-adsorption from solution phase-applications of adsorption.

Colloids; Classification of colloids: Classification based on physical state of dispersed phase and dispersion medium- classification based on nature of interaction between dispersed phase and dispersion medium- classification based on type of particles of the dispersed phase- multi molecular, macromolecular and associated colloids- cleansing action of soaps-preparation of colloids-purification of colloidal solutions- properties of colloidal solutions: Tyndall effect, colour, Brownian movement-charge on colloidal particles, electrophoresis. Coagulation or precipitation, Coagulation of lyophilic sols, protection of colloids, Colloids around us-applications of colloids.

17) p-BLOCK ELEMENTS:

GROUP-15 ELEMENTS : Introduction - Occurrence- electronic configuration, atomic and ionic radii, ionisation enthalpy, electronegativity, physical and chemical properties; Dinitrogen-preparation, properties and uses; Compounds of nitrogen-preparation, properties and uses of ammonia; Oxides of nitrogen; preparation, properties and uses of nitric acid; Phosphorous-allotropic forms.

GROUP-16 ELEMENTS: Introduction - Occurrence- electronic configuration, atomic and ionic radii, ionisation enthalpy, electron gain enthalpy, electronegativity, physical and chemical properties; Dioxygen-preparation, properties and uses; Simple oxides; Ozone-preparation, properties, structure and uses; Sulphur-allotropic forms; Sulphur dioxide-preparation, properties and uses; Oxoacids of sulphur; Sulphuric acid-Properties and uses only.

GROUP-17 ELEMENTS: Introduction - Occurrence, electronic configuration, atomic and ionic radii, ionisation enthalpy, electron gain enthalpy, electronegativity, physical and



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chemical properties; Chlorine- preparation, properties and uses; Hydrogen chloride- preparation, properties and uses; Oxoacids of halogens; Interhalogen compounds.

GROUP-18 ELEMENTS : Introduction - Occurrence, electronic configuration, ionization enthalpy, atomic radii, electron gain enthalpy, physical and chemical properties (a) Xenon-fluorine compounds- XeF_2 , XeF_4 and XeF_6 -preparation, hydrolysis and formation of fluoro anions-structures of XeF_2 , XeF_4 and XeF_6 (b) Xenon-oxygen compounds XeO_3 and XeOF_4 - their formation, structures and uses.

18) d AND f BLOCK ELEMENTS & COORDINATION COMPOUNDS:

d AND f BLOCK ELEMENTS : Position in the periodic table; Electronic configuration of the d-block elements; General properties of the transition elements (d-block) -physical properties, variation in atomic and ionic sizes of transition series, ionisation enthalpies, oxidation states, trends in the M^{2+}/M and M^{3+}/M^{2+} standard electrode potentials, trends in stability of higher oxidation states, chemical reactivity and E^\ominus values, magnetic properties, formation of coloured ions, formation of complex compounds, catalytic properties, formation of interstitial compounds, alloy formation.

COORDINATION COMPOUNDS: Werner's theory of coordination compounds; Definitions of some terms used in coordination compounds; Nomenclature of coordination compounds-IUPAC nomenclature; Isomerism in coordination compounds- (a) Stereo isomerism-Geometrical and optical isomerism (b) Structural isomerism-linkage, coordination, ionisation and hydrate isomerism; Bonding in coordination compounds. (a) Valence bond theory - magnetic properties of coordination compounds-limitations of valence bond theory (b) Crystal field theory (i) Crystal field splitting in octahedral and tetrahedral coordination entities (ii) Colour in coordination compounds-limitations of crystal field theory; Bonding in metal carbonyls; Stability of coordination compounds; Importance and applications of coordination compounds.

19) BIOMOLECULES:

Carbohydrates - Classification of carbohydrates-Monosaccharides: preparation of glucose from sucrose and starch- Properties and structure of glucose- D, L and (+), (-) configurations of glucose- Structure of fructose.

Amino acids: Natural amino acids-classification of amino acids - structures and D and L forms-Zwitterions;

Proteins: Structures, classification, fibrous and globular- primary, secondary, tertiary and quaternary structures of proteins- Denaturation of proteins.

Nucleic acids: Chemical composition of nucleic acids, structures of nucleic acids, DNA finger printing, biological functions of nucleic acids.

20) HALOALKANES AND HALOARENES:

Classification and nomenclature; Nature of C-X bond; Methods of preparation: Alkyl halides and aryl halides-from alcohols, from hydrocarbons (a) by free radical halogenation (b) by electrophilic substitution (c) by replacement of diazonium group (Sandmeyer reaction) (d) by the addition of hydrogen halides and halogens to alkenes-by halogen exchange (Finkelstein reaction); Physical properties-melting and boiling points, density and solubility; Chemical reactions: Reactions of haloalkanes (i) Nucleophilic substitution reactions (a) SN^2 mechanism (b) SN^1 mechanism (c) stereochemical aspects of nucleophilic substitution reactions-optical activity (ii) Elimination reactions (iii) Reaction with metals-Reactions of halo arenes: (i) Nucleophilic substitution (ii) Electrophilic substitution and (iii) Reaction with metals.



Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

21) ORGANIC COMPOUNDS CONTAINING C, H AND O (Alcohols, Phenols, Ethers, Aldehydes, Ketones and Carboxylic acids):

ALCOHOLS, PHENOLS AND ETHERS

Alcohols, phenols and ethers -classification; Nomenclature: (a)Alcohols, (b)phenols and (c) ethers; Structures of hydroxy and ether functional groups; Methods of preparation: **Alcohols** from alkenes and carbonyl compounds (reduction and reaction with Grignard reagents); **Phenols** from haloarenes, benzene sulphonic acid, diazonium salts, cumene;

Physical properties of alcohols and phenols; Chemical reactions of alcohols and phenols (i) Reactions involving cleavage of O-H bond-Acidity of alcohols and phenols, esterification (ii) Reactions involving cleavage of C-O bond- reactions with HX, PX₃, dehydration and oxidation (iii) Reactions of phenols- electrophilic aromatic substitution, Kolbe's reaction, Reimer - Tiemann reaction, reaction with zinc dust, oxidation.

Ethers-Methods of preparation: By dehydration of alcohols, Williamson synthesis-Physical properties-Chemical reactions: Cleavage of C-O bond and electrophilic substitution of aromatic ethers.

ALDEHYDES AND KETONES

Nomenclature and structure of carbonyl group; Preparation of aldehydes and ketones- (1) by oxidation of alcohols (2) by dehydrogenation of alcohols (3) from hydrocarbons -Preparation of aldehydes (1) from acyl chlorides (2) from nitriles and esters(3) from hydrocarbons-Preparation of ketones (1) from acyl chlorides (2)from nitriles (3)from benzene or substituted benzenes; Physical properties of aldehydes and ketones; Chemical reactions of aldehydes and ketones-nucleophilic addition, reduction, oxidation, reactions due to alpha hydrogen and other reactions (Cannizzaro reaction, electrophilic substitution reaction); Uses of aldehydes and ketones.

CARBOXYLIC ACIDS

Nomenclature and structure of carboxyl group; Methods of preparation of carboxylic acids (1)from primary alcohols and aldehydes (2) from alkyl benzenes(3)from nitriles and amides (4)from Grignard reagents (5) from acyl halides and anhydrides (6) from esters; Physical properties; Chemical reactions: (i) Reactions involving cleavage of O-H bond-acidity, reactions with metals and alkalies (ii) Reactions involving cleavage of C-OH bond-formation of anhydride, reactions with PCl₅, PCl₃, SOCl₂, esterification and reaction with ammonia (iii) Reactions involving-COOH group-reduction, decarboxylation (iv) Substitution reactions in the hydrocarbon part - halogenation and ring substitution; Uses of carboxylic acids.

22) ORGANIC COMPOUNDS CONTAINING NITROGEN:

AMINES

Structure of amines; Classification; Nomenclature; Preparation of amines: reduction of nitro compounds, ammonolysis of alkyl halides, reduction of nitriles, reduction of amides, Gabriel phthalimide synthesis and Hoffmann bromamide degradation reaction; Physical properties; Chemical reactions: basic character of amines, alkylation, acylation, carbyl amine reaction, reaction with nitrous acid, reaction with aryl sulphonyl chloride, electrophilic substitution of aromatic amines-bromination, nitration and sulphonation.

Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

MODEL QUESTIONS – MATHEMATICS

- 1) I) The coefficient of x^5 in $(1 - x - x^2 + x^3)^6$ is 20.
 II) If p and q are the coefficients of x^4 in $(1 + x)^{2n}$ and $(1 + x)^{2n-1}$ respectively then $2p=q$.
 Which of the above statements is (are) true.

- 1) only I 2) Only II 3) Both I and II 4) Neither I nor II

- 2). Assertion (A): If $A+B=45^\circ$ then $(1+\tan A)(1+\tan B)=2$.
 Reason(R): $\tan 22\frac{1}{2}^\circ = \sqrt{2} - 1$.

- 1) Both A & R are True and R is the correct explanation of A.
 2) Both A & R are True and R is not correct explanation of A.
 3) A is True but R is False.
 4) R is True but A is False.

- 3) Arrange the following statements in ascending order of their results

- A) The order of $\left(\frac{d^4y}{dx^4} + \frac{d^2y}{dx^2}\right)^{3/2} = a\left(\frac{d^3y}{dx^3}\right)$
 B) The degree of $\left(\frac{dy}{dx} + \frac{d^2y}{dx^2}\right)^{5/4} = a\left(\frac{d^3y}{dx^3}\right)^{2/3}$
 C) The degree of $y = \left(1 + \left(\frac{dy}{dx}\right)^2\right)^{3/2} \left(\frac{d^2y}{dx^2}\right)$
 D) The order of $x^3 + \left(\frac{d^3y}{dx^3}\right)^2 + 2x^2 \left(\frac{d^2y}{dx^2}\right) - 3y = x^2$

- 1) D,C,B,A 2) D,C,A,B 3) C,D,B,A 4) C,D,A,B

- 4) A and B are two independent events of a sample space such that $P(A)=0.2$, $P(B)=0.5$.

- | | |
|------------------|----------|
| List I | List II |
| A) $P(B/A)$ | I) 0.2 |
| B) $P(A/B)$ | II) 0.1 |
| C) $P(A \cap B)$ | III) 0.3 |
| D) $P(A \cup B)$ | IV) 0.6 |
| | V) 0.5 |

The correct match is

- | | | | | |
|----|-----|---|-----|----|
| 1) | A | B | C | D |
| | IV | V | III | I |
| 2) | A | B | C | D |
| | V | I | II | IV |
| 3) | A | B | C | D |
| | III | I | II | IV |
| 4) | A | B | C | D |
| | II | I | IV | V |

Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

5) The line $ax + by + c = 0$ is a normal to the circle $x^2 + y^2 + 4x + 6y + 8 = 0$ if and only if

(1) $2a + 3b = c$ (2) $3a + 2b = c$ (3) $2a + 3b + c = 0$ (4) $3a + 2b + c = 0$

6) $\begin{vmatrix} al & bm & cn \\ l^2 & m^2 & n^2 \\ 1 & 1 & 1 \end{vmatrix} = \Delta_1$ and $\begin{vmatrix} a & b & c \\ l & m & n \\ mn & ln & lm \end{vmatrix} = \Delta_2$ then

(1) $\Delta_1 = \Delta_2$

(2) $\Delta_1 = 2\Delta_2$

(3) $2\Delta_1 = \Delta_2$

(4) $\Delta_1 + \Delta_2 = 0$

7) If \vec{a} is a non-zero vector and \vec{b}, \vec{c} are two vectors such that $\vec{a} \times \vec{b} = \vec{a} \times \vec{c}$ and $\vec{a} \cdot \vec{b} = \vec{a} \cdot \vec{c}$ then

1) $\vec{b} - \vec{c}$ is collinear with \vec{a}

2) $\vec{b} - \vec{c}$ is perpendicular with \vec{a}

3) $\vec{b} = \vec{c}$

4) $\vec{b} \neq \vec{c}$

8) If $L_1 : 2x + 3y - 20 = 0$, $L_2 : 2x + 3y - 14 = 0$, then the straight line represented by $a(2x + 3y - 20) + b(2x + 3y - 14) = 0$ is

1) Parallel to $L_1 = 0$ and $L_2 = 0$.

2) Perpendicular to $L_1 = 0$ and parallel to $L_2 = 0$

3) Perpendicular to $L_1 = 0$ and $L_2 = 0$

4) Parallel to $L_1 = 0$ and Perpendicular to $L_2 = 0$

9) If $f(x) = \log\left(\frac{1+2x}{1-2x}\right)$, then $x =$

1) $\frac{e^{f(x)} - 1}{2(e^{f(x)} + 1)}$

2) $\frac{2(e^{f(x)} - 1)}{(e^{f(x)} + 1)}$

3) $\frac{e^{f(x)} - 1}{(e^{f(x)} + 1)}$

4) $\frac{e^{f(2x)} - 1}{(e^{f(2x)} + 1)}$

10) If $\tan\theta = \frac{b}{a}$ then $a \cos 2\theta - b \sin 2\theta =$

1) a

2) b

3) $2a$

4) $2b$



MODEL QUESTIONS – PHYSICS

1. A particle performs simple harmonic motion with amplitude A and time period T . The mean velocity of the particle over the time interval during which it travels a distance of $A/2$ starting from extreme position.

- 1) A/T 2) $2A/T$ 3) $3A/T$ 4) $A/2T$

2. When a battery connected across a resistor of 16Ω , the voltage across the resistor is $12V$. When same battery is connected across a resistor of 10Ω voltage across it is $11V$. The internal resistance of the battery

- 1) $10/7\Omega$ 2) $20/7\Omega$ 3) $25/7\Omega$ 4) $30/7\Omega$

3. **Assertion (A):** A rocket works on the principle of conservation of linear momentum.

Reason (R): Whenever there is change in momentum of one body, the same change occurs in the momentum of the second body of the same system but in the opposite direction.

- 1) A is true & R is true and correct explanation
2) A is true & R is true and not correct explanation
3) A is true & R is false
4) A is false & R is true

4. **Statement(A):** A particle can have zero displacement and non zero average velocity.

Statement (B): A particle can have zero acceleration and non zero velocity

Statement (C): A particle can have zero velocity and non-zero acceleration.

- 1) A,B,C True 2) A, B True, C False 3) B,C True, A False 4) A,B,C False.

5. Match the following

In the experimental study of photoelectric effect:

Column-I

- A.** Intensity of incident light changes
B. Frequency of incident light changes
C. Target material changes

Column-II

- I.** Maximum K.E of photoelectrons changes
II. Stopping potential changes
III. Saturation current changes.

1. A-III B-I,II C-I,II
2. A-II B-I,III C-I,II
3. A-III B-III,II C-I,II
4. A-I B-I,II C-I,II

Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

MODEL QUESTIONS – CHEMISTRY

- Half life of a first order reaction is 15min. The time required for completion of 87.5% of the reaction is
(1) 15 min (2) 30 min (3) 60 min (4) 45 min
- Arrange the following in increasing order of their basic strength.
I) $C_2H_5NH_2$ II) $C_6H_5NH_2$ III) $C_6H_5CH_2NH_2$ IV) $(C_2H_5)_2NH$
(1) $I < II < IV < III$ (2) $II < III < I < IV$
(3) $I < IV < III < II$ (4) $IV < I < III < II$
- Which of the following has least electron gain enthalpy?
(1) Sulphur (2) Chlorine (3) Flourine (4) Oxygen
- Assertion (A): H_2O is liquid and H_2S is gas at room temperature
Reason (R) : Molecules of H_2O are highly associated through hydrogen bonding
The correct answer is:
(1) Both (A) and (R) are true and (R) is the correct explanation of (A)
(2) Both (A) and (R) are true and (R) is not the correct explanation of (A)
(3) (A) is true but (R) is not true
(4) (A) is not true but (R) is true

5. Match the following:

LIST I (Crystal System)

- (A) Cubic
(B) Hexagonal
(C) Monoclinic
(D) Triclinic

LIST II (Axial Angle)

- (I) $\alpha = \beta = 90^\circ; \gamma = 120^\circ$
(II) $\alpha \neq \beta \neq \gamma \neq 90^\circ$
(III) $\alpha = \beta = \gamma \neq 90^\circ$
(IV) $\alpha = \gamma = 90^\circ; \beta \neq 90^\circ$
(V) $\alpha = \beta = \gamma = 90^\circ$

The correct match is:

- | | | | | |
|-----|----|----|-----|-----|
| (1) | A | B | C | D |
| | V | I | III | IV |
| (2) | A | B | C | D |
| | IV | II | V | I |
| (3) | A | B | C | D |
| | V | I | IV | II |
| (4) | A | B | C | D |
| | IV | II | V | III |



Deleted topics of Mathematics

The following topics are deleted (30%) in the content of MATHEMATICS Text Book Published by Telugu Academy in the Intermediate First Year (IA & IB) for the Academic Year 2020-21 and Second year (IIA & IIB) for the Academic Year 2021-22 by the Telangana Board of Intermediate Education (TSBIE). The same 30% content is deleted for the TS EAMCET-2022 Examination.

Following are the 30% Deleted Topics in the First-Year (IA) Intermediate Mathematics Syllabus.

CHAPTER 1: Functions

1.2 Inverse functions and theorems

CHAPTER 2: Mathematical induction

Complete chapter

CHAPTER 3: Matrices

3.4.8 Properties of Determinants related Problems and Examples

3.6 Consistency and inconsistency of system of simultaneous linear equations
Gauss – Jordan Method

3.7.7 onwards

CHAPTER 5: Product of Vectors

5.11 Vector equation of a plane – different forms, skew – lines, shortest distance
– plane, condition for co-planarity etc.

5.12 Vector triple product – results

CHAPTER 7: Trigonometric Equations

Complete chapter

CHAPTER 8: Inverse Trigonometric Functions

Complete chapter

CHAPTER 9: Hyperbolic functions

9.2 Definition of Inverse hyperbolic functions and graphs

Following are the 30% Deleted Topics in the First-Year (IB) Intermediate Mathematics Syllabus.

CHAPTER 4: Pair of straight lines

4.3 Pair of bisectors of angles and related problems

4.5 Condition of pair of parallel lines and distance between them and point of intersection of pair of lines, Exercise 4(b)

CHAPTER 7: Plane

Exercise 7(a) Section II & III related Examples

Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

CHAPTER 8: Limits and Continuity

8.4 Continuity

CHAPTER 9: Differentiation

9.3 Problems of inverse trigonometric functions Exercise 9(c) Section III.

9.5 Second order derivatives

CHAPTER 10: Applications of Derivatives

10.4 Lengths of tangent, normal, sub-tangent and sub-normal

10.6 Derivative as a rate of change

10.7 Rolle's, Lagrange's, mean value theorem

10.8 Increasing and decreasing functions and related Problems and Examples

Following are the 30% Deleted Topics in the Second-Year (IIA) Intermediate Mathematics Syllabus.

CHAPTER 1: Complex Numbers

1.3 Modulus and amplitudes of complex number-illustrations.

1.4 Geometrical and polar representation of a complex number in Argand plane – Argand Diagram.

CHAPTER 2: De Moivre's Theorem

Exercise 2(b) section II and section III.

CHAPTER 3: Quadratic expressions

3.3 Quadratic inequalities.

CHAPTER 5: Permutations and Combinations

5.3 Permutations when repetitions are allowed.

5.4 Circular permutations.

5.5 Permutations with constraint repetitions.

Exercise 5(e) Section III.

CHAPTER 6: Binomial Theorem

Exercise 6(a) Section II 5th problem onwards and related examples.

Exercise 6(b) Section II and related examples.

Exercise 6(c).

CHAPTER 7: Partial Fractions

Exercise 7(d).

CHAPTER 8: Measures of Dispersion

8.2.2 Mean Deviation for grouped data.

Exercise 8(a) Section I Problem 3 onwards

8.3 Variance and standard deviation of ungrouped / grouped data.

8.4 Coefficient of Variation and analysis of frequency distributions with equal means but different variances.

CHAPTER 9: Probability

9.3.9 Baye's Theorem and problems on Baye's Theorem.



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Following are the 30% Deleted Topics in the Second-Year (IIB) Intermediate Mathematics Syllabus.

CHAPTER 3: Parabola

3.2 Equation of tangent and normal at a point on the parabola.

CHAPTER 4: Ellipse

4.2 Equation of tangent and normal at a point on the ellipse.

CHAPTER 5: Hyperbola

5.2 Equation of tangent and normal at a point on the hyperbola Exercise 5(a) Section II onwards and related examples.

CHAPTER 6: Integration

6.2(b) Integration by parts – Integration of exponential, logarithmic and inverse trigonometric functions.

6.3 Integration-partial fraction method.

6.4 Reduction formulae.

CHAPTER 7: Definite Integrals

Problems and Examples related to inverse trigonometric functions

Exercise-7(b) Section II (8 to 15).

7.5 Reduction formula.

7.6 Application of definite integrals to areas.

CHAPTER 8: Differential Equations

8.2(b) Homogeneous Differential Equations.

8.2(c) Non-Homogeneous Differential Equations.

8.2(d) Linear Differential Equations.



Deleted topics of Physics

The following topics are Deleted (30%) in the content of PHYSICS Text Book published by Telugu Academy in the Intermediate First Year for the Academic Year 2020-21 and Second Year for the Academic Year 2021-22 by Telangana State Board of Intermediate Education (TSBIE). The same 30% content is deleted for the TS EAMCET-2022 Examination.

Following are the 30% Deleted Topics in the First-Year Intermediate Physics Syllabus.

CHAPTER-1 : PHYSICAL WORLD

- 1.2 Scope and excitement of Physics.
- 1.3 Physics, technology and society.
- 1.5 Nature of physical laws.

CHAPTER-5: LAWS OF MOTION

- 5.2 Aristotle's fallacy.
- 5.3 The law of inertia.
- 5.4 Newton's first law of motion.
- 5.5 Newton's second law of motion.
- 5.6 Newton's third law of motion,

(These topics are deleted, however they must be recapitulated as a pre-requisite to deal with the remaining topics of the chapter.)

CHAPTER-7 : SYSTEMS OF PARTICLES AND ROTATIONAL MOTION

- 7.10 Theorems of perpendicular and parallel axes.

CHAPTER-9: GRAVITATION

- 9.2 Kepler's laws.
- 9.4 The gravitational constant. (Despite the topic is deleted, the value of G should be mentioned to the student)
- 9.5 Acceleration due to gravity of the earth.

CHAPTER-10: MECHANICAL PROPERTIES OF SOLIDS:

- 10.6.2 Determination of Young's Modulus of the Material of a Wire.

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- 10.6.3 Shear Modulus.
- 10.6.5 Poisson's Ratio.
- 10.6.6 Elastic potential energy in a stretched wire.
- 10.7 Applications of elastic behaviour of materials

CHAPTER-12: THERMAL PROPERTIES OF MATTER:

- 12.9 Heat transfer.
- 12.9.1 Conduction.
- 12.9.2 Convection.
- 12.9.3 Radiation.

(These topics are deleted, however they must be recapitulated as a pre-requisite to deal with the remaining topics of the chapter.)

CHAPTER-13: THERMODYNAMICS:

- 13.9 Heat engines.
- 13.10 Refrigerators and heat pumps.
- 13.13 Carnot engine

Following are the 30% Deleted Topics in the Second-Year Intermediate Physics Syllabus.

CHAPTER-1: WAVES

- 1.8 Doppler Effect

CHAPTER-2: RAY OPTICS AND OPTICAL INSTRUMENTS

- 2.2 Reflection of light by Spherical Mirrors
- 2.8.2 Scattering of light

CHAPTER-3: WAVE OPTICS

- 3.6.3 Resolving power of optical instruments
- 3.7 Polarization

CHAPTER-4: ELECTRIC CHARGES AND FIELDS

- 4.15.3 Field due to a uniformly charged thin spherical shell

CHAPTER-5: ELECTROSTATIC POTENTIAL AND CAPACITANCE

- 5.16 Van de Graaf generator

CHAPTER-6: CURRENT ELECTRICITY

- 6.7 Resistivity of various Materials
- 6.10 Combination of resistors-series and parallel



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CHAPTER-7: MOVING CHARGES AND MAGNETISM

7.4 Motion in combined electric and magnetic fields

CHAPTER-8: MAGNETISM AND MATTER

8.2.2 Bar Magnet as a equivalent solenoid

8.2.3 The dipole in a uniform magnetic field

8.6 Magnetic properties of materials

8.7 Permanent magnets and electromagnets

CHAPTER-10: ALTERNATING CURRENT

10.7 Power in AC Circuit: The Power Factor

CHAPTER-11: ELECTRO MAGNETIC WAVES

11.2 Displacement Current

CHAPTER-12: DUAL NATURE OF RADIATION AND MATTER

12.9 Davisson and Germer Experiment

CHAPTER-14: NUCLEI

14.4.2 Nuclear Binding energy

14.6 Radioactivity

**CHAPTER-15: SEMICONDUCTOR ELECTRONICS: MATERIALS, DEVICES
AND SIMPLE CIRCUITS**

15.8.1 Zener diode

15.9.3 Transistor as a device

15.9.4 Transistor as an amplifier (CE configuration)

15.9.5 Feedback amplifier and transistor oscillator



Deleted topics of Chemistry

The following topics are the deleted (30%) in the content of CHEMISTRY Text book published by Telugu Academy in the Intermediate First year for the academic year 2020-21 and second year for the academic year 2021-22 by Telangana State Board of Intermediate Education (TSBIE). The same 30% content is deleted for the TS EAMCET-2022 examination.

Following are the 30% deleted topics in the First-Year Intermediate Chemistry syllabus

Chapter 1: ATOMIC STRUCTURE:

- 1.1 Sub- atomic particles
- 1.2 Atomic models – Thomson’s Model; Rutherford’s Nuclear model of atom, Drawbacks
- 1.6 Quantum mechanical model of an atom. Important features of Quantum mechanical model of atom
 - 1.6.2 Shapes of atomic orbitals

Chapter 2: CLASSIFICATION OF ELEMENTS AND PERIODICITY IN PROPERTIES:

- 2.1 Need to classify elements
- 2.2 Genesis of periodic classification

Chapter 4: STATES OF MATTER: GASES AND LIQUIDS:

- 4.10 Distribution of molecular speeds - rms, average and most probable speeds-Kinetic energy of gas molecules
- 4.12 Liquefaction of gases
- 4.13 Liquid State - Properties of Liquids in terms of Inter molecular interactions - Vapour pressure, Viscosity and Surface tension (Qualitative idea only. No mathematical derivation).

Chapter 5: STOICHIOMETRY:

- 5.2 Laws of Chemical Combinations - Law of Conservation of Mass, Law of Definite Proportions, Law of Multiple Proportions, Gay Lussac’s Law of Gaseous Volumes, Dalton’s Atomic Theory, Avogadro’s Law
- 5.6 Methods of expressing concentrations of solutions-mass percent, mole fraction, molarity, molality and normality;
- 5.11 Redox reactions in Titrimetry.

Chapter 6: THERMODYNAMICS:

- 6.2.4 Heat capacity
- 6.2.5 The relationship between C_p and C_v
- 6.2.7 Gibbs Energy change and equilibrium

Chapter 7: CHEMICAL EQUILIBRIUM AND ACIDS-BASES:

- 7.7 Relationship between Equilibrium constant K, reaction quotient Q and Gibbs energy G;

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- 7.11 Ionisation of Acids and Bases -Ionisation constant of water and its ionic product- pH scale-ionisation constants of weak acids-ionisation of weak bases-relation between K_a and K_b -Di and poly basic acids and di and poly acidic Bases-Factors affecting acid strength-Common ion effect in the ionization of acids and bases-Hydrolysis of salts and pH of their solutions.
- 7.12 Henderson equation derivation

Chapter 8: HYDROGEN AND ITS COMPOUNDS:

- 8.3 Preparation of Dihydrogen
- 8.4 properties of Dihydrogen
- 8.7 Hydrogen peroxide: Preparation; Physical properties; structure and chemical properties; storage and uses

Chapter 9: s - BLOCK ELEMENTS (ALKALI AND ALKALINE EARTH METALS)

- 9.4 Some important compounds of Sodium: Sodium Carbonate; Sodium Chloride; Sodium Hydroxide; Sodium hydrogen carbonate
- 9.5 Biological importance of Sodium and Potassium.
- 9.9 Some important compounds of calcium: Preparation and uses of Calcium Oxide; Calcium Hydroxide; Calcium Carbonate; Plaster of Paris; Cement
- 9.10 Biological importance of Calcium and Magnesium.

Chapter 10: p- BLOCK ELEMENTS GROUP 13 (BORON FAMILY):

- 10.3 Some important compounds of boron - Borax, Ortho boric acid, diborane
- 10.4 Uses of boron, aluminium and their compounds.

Chapter 11: p-BLOCK ELEMENTS - GROUP 14 (CARBON FAMILY):

- 11.4 Some important compounds of carbon and silicon – carbon monoxide, carbon dioxide, Silica, silicones, silicates and zeolites.

Chapter 12: ENVIRONMENTAL CHEMISTRY: (ENTIRE CHAPTER)

Definition of terms: Air, Water and Soil Pollutions; Environmental Pollution; Atmospheric pollution; Tropospheric Pollution; Gaseous Air Pollutants (Oxides of Sulphur; Oxides of Nitrogen; Hydrocarbons; Oxides of Carbon (CO, CO₂)). Global warming and Green house effect; Acid rain- Particulate Pollutants- Smog; Stratospheric Pollution: Formation and breakdown of Ozone- Ozone hole- effects of depletion of the Ozone Layer; Water Pollution: Causes of Water Pollution; International standards for drinking water; Soil Pollution: Pesticides, Industrial Wastes; Strategies to control environmental pollution- waste Management- collection and disposal; Green Chemistry: Green chemistry in day-to-day life; Dry cleaning of clothes; Bleaching of paper; Synthesis of chemicals

Chapter 13: ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES, TECHNIQUES AND HYDROCARBONS

- 13.8 Methods of purification of organic compounds;
- 13.9 Qualitative elemental analysis of organic compounds;
- 13.10 Quantitative elemental analysis of organic compounds.
- 13.12.2 Alkanes properties



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1. Free radical mechanism of halogenations,
2. Combustion
7. Pyrolysis

Following are the 30% deleted topics in the Second-Year Intermediate Chemistry Syllabus

Chapter 1: Solid state

- 1.11 Electrical properties
- 1.12 Magnetic properties (Band theory of metals, conductors, semiconductors and insulators and n and p type semiconductors)

Chapter 2: Solutions

- 2.7 Abnormal molar masses-van't Hoff factor.

Chapter 3: Electrochemistry and Chemical Kinetics

Electrochemistry

- 3.1 Electrochemical cells
- 3.2 Galvanic cells
- 3.5 Electrolytic cells
- 3.6 Batteries: primary and secondary batteries
- 3.7 Fuel cells
- 3.8 Corrosion of metals-Hydrogen economy

Chemical kinetics

- 3.14 Collision theory of chemical reaction rates.

Chapter 4: Surface Chemistry

- 4.2 catalysis, homogenous and heterogenous, Adsorption theory of heterogenous catalysis, activity and selectivity of solid catalysis, Shape-selective catalysis by zeolites, Enzyme catalysis, Catalysts in industry.
- 4.5 Emulsions – types of emulsions

Chapter 5 : General Principles of Metallurgy

Entire chapter is deleted

Chapter 6 : P-block Elements

Group-15 Elements

- 6.7 Phosphine-preparation, properties and uses
- 6.8 Phosphorous halides
- 6.9 Oxoacids of phosphorous

Group -16 Elements

- 6.17 Sulphuric acid-manufacture.

Chapter 7: d and f Block Elements & Coordination Compounds

d and f Block Elements

- 7.4 Some important compounds of transition elements
- 7.5 Inner transition elements
- 7.6 Actinoids
- 7.7 Some applications of d and f block elements.



Syllabus for TS EAMCET 2022-E Stream (Engineering Stream)

Chapter 8 : Polymers

Entire chapter is deleted

Chapter 9 : Biomolecules

9.1 Disaccharides (sucrose, lactose, maltose), polysaccharides(starch, cellulose, glycogen), importance of carbohydrates)

9.3 Enzymes: Enzymes, mechanism of enzyme action

9.4 Vitamins

9.6 Hormones.

Chapter 10 : Chemistry in Everyday life

Entire chapter is deleted

Chapter 11 : Halo Alkanes and Halo Arenes

11.6 Polyhalogen compounds.

Chapter 12 : Organic Compounds Containing C, H and O

12.7 Some Commercially important alcohols (uses with special reference to methanol and ethanol)

Chapter 13 : Organic Compounds Containing Nitrogen

II. Diazonium salts

13.7 Methods of preparation of diazonium salts

13.8 Physical properties

13.9 Chemical reactions

13.10 Importance of diazonium salts in synthesis of aromatic compounds

III. Cyanides and Isocyanides

13.11 Structure and nomenclature of cyanides and isocyanides

13.12 Preparation, physical properties and chemical reactions of cyanides and isocyanides.

